



# MODELLING OF MASS TRANSFER THROUGH THE BARRIER ENVELOPES OF VIPS

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Lionel FLANDIN, Bernard YRIEIX

IVIS 2015 – Nanjing – September 19th and 20th, 2015



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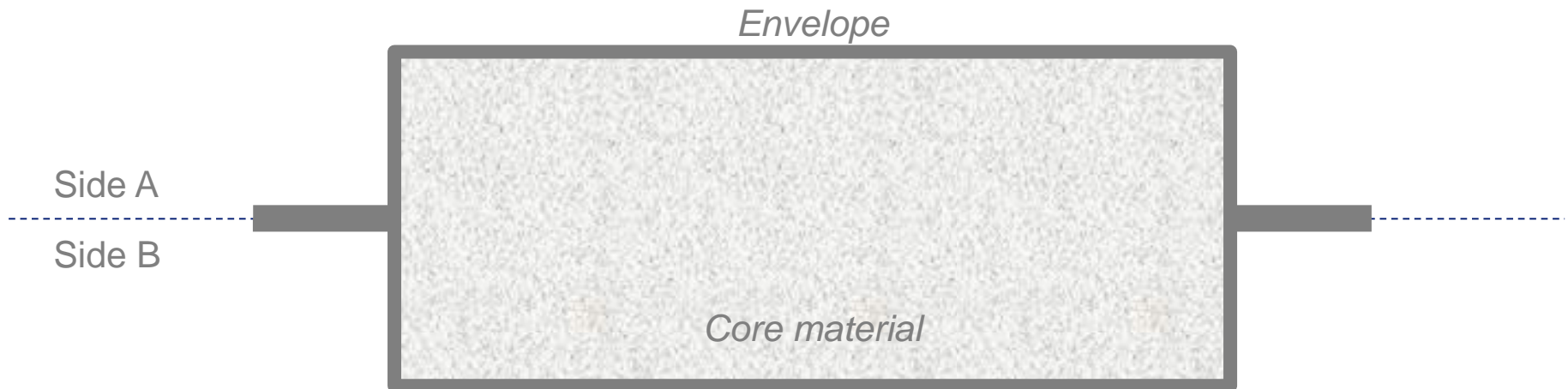
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2. PURPOSES OF THE STUDY
3. MODELLING APPROACH
4. SOME RESULTS WITH PET M1F
5. CONCLUSION AND OUTLOOK



# INTRODUCTION

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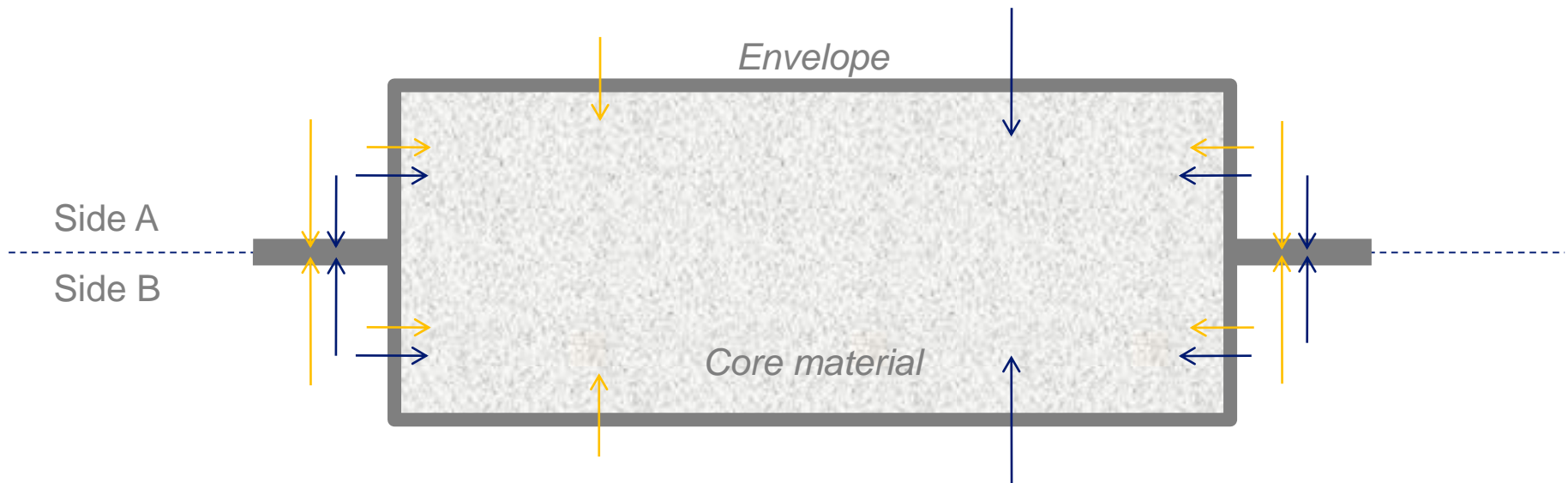
- Hygrothermal behaviour of VIPs



# INTRODUCTION

→ Dry air  
→ Water vapour

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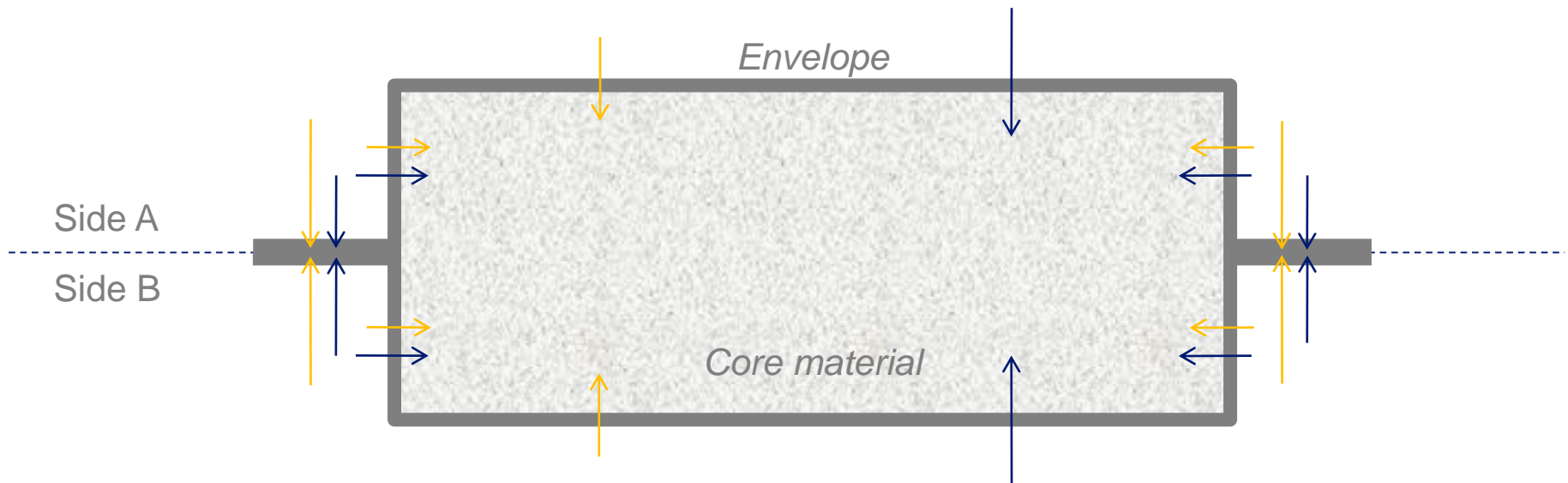


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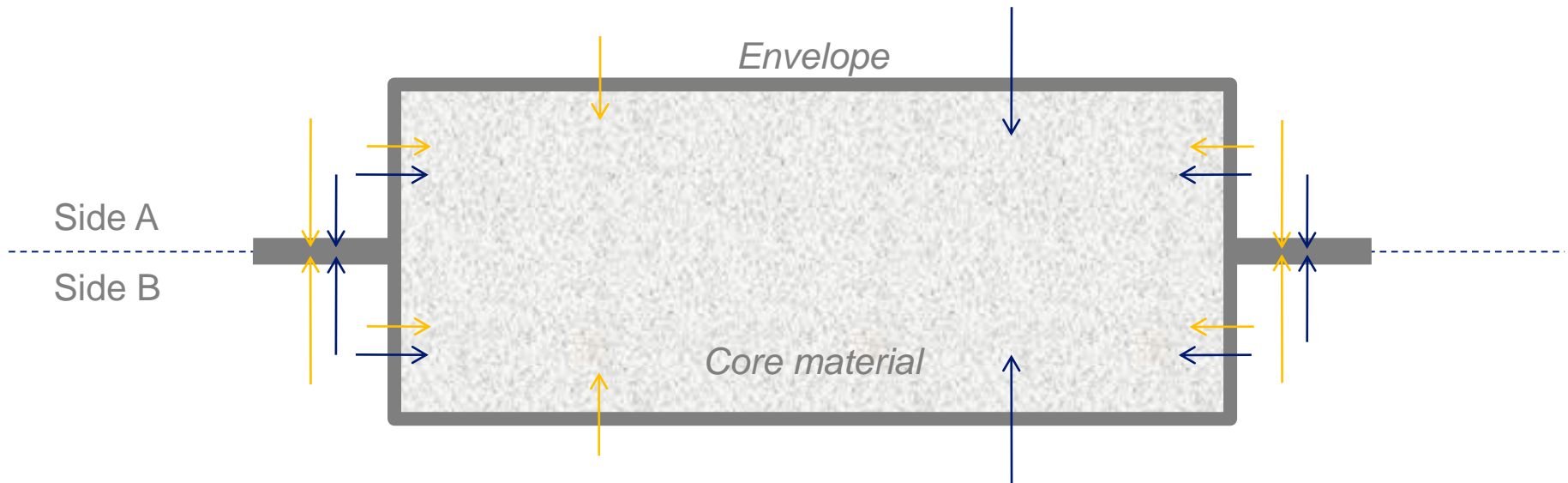
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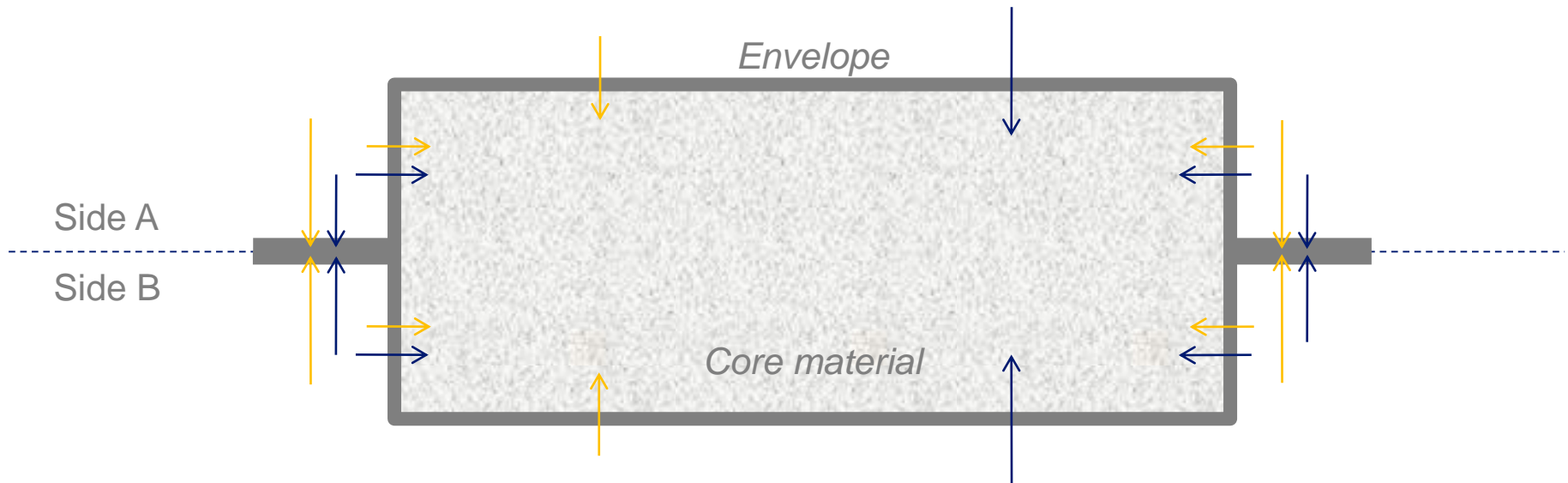
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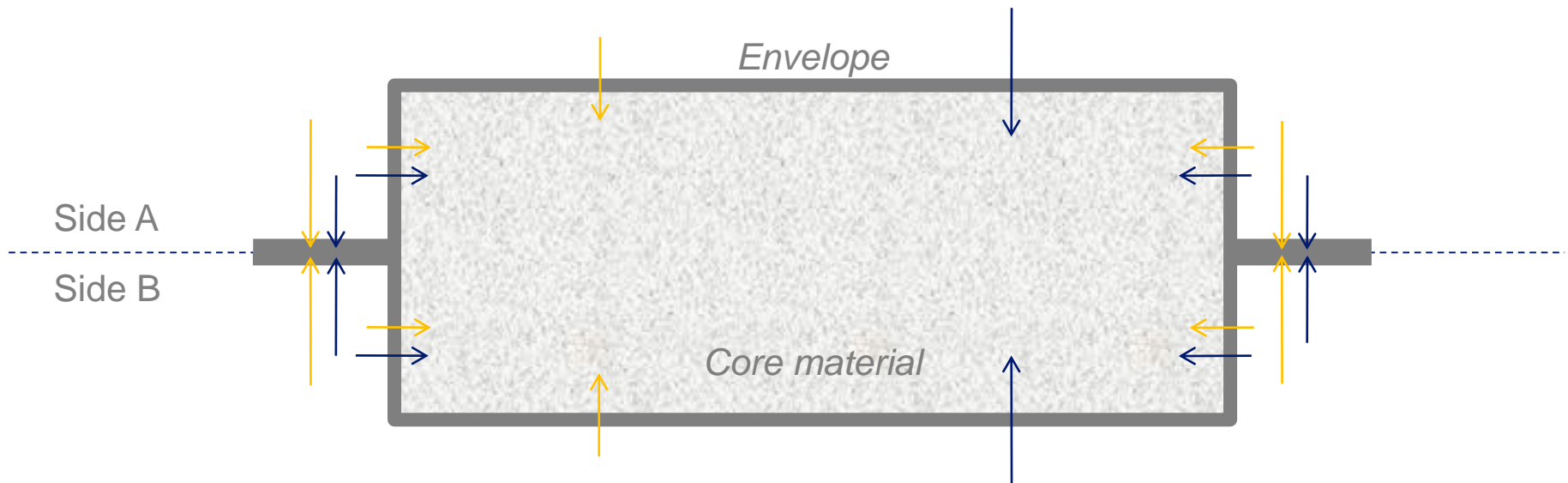
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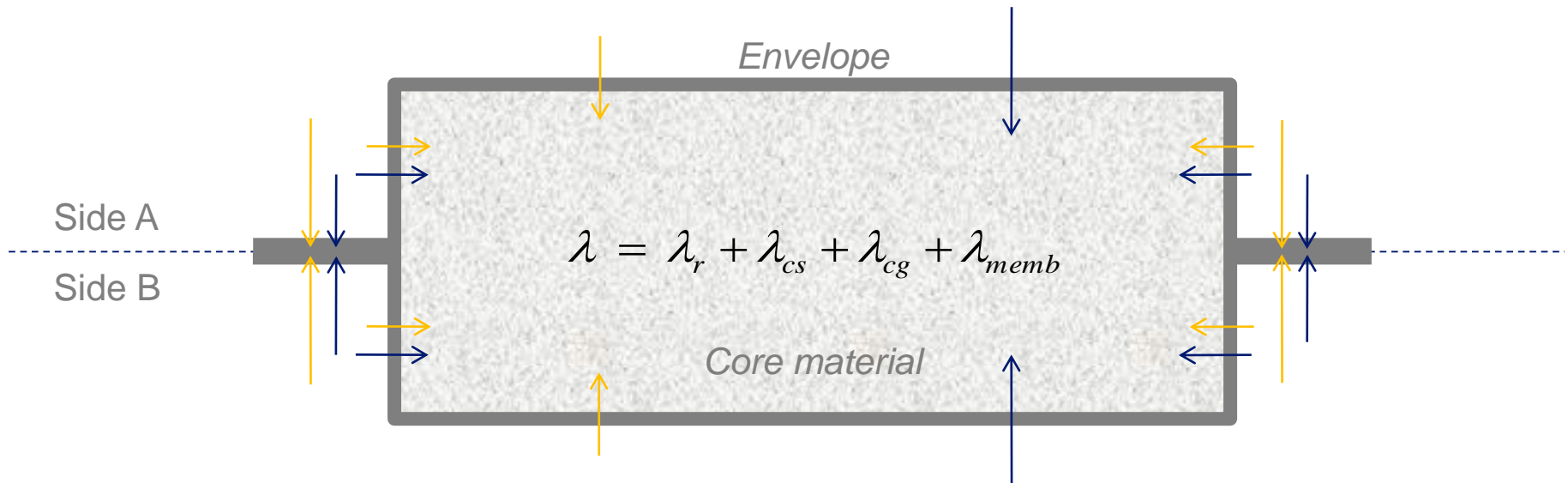
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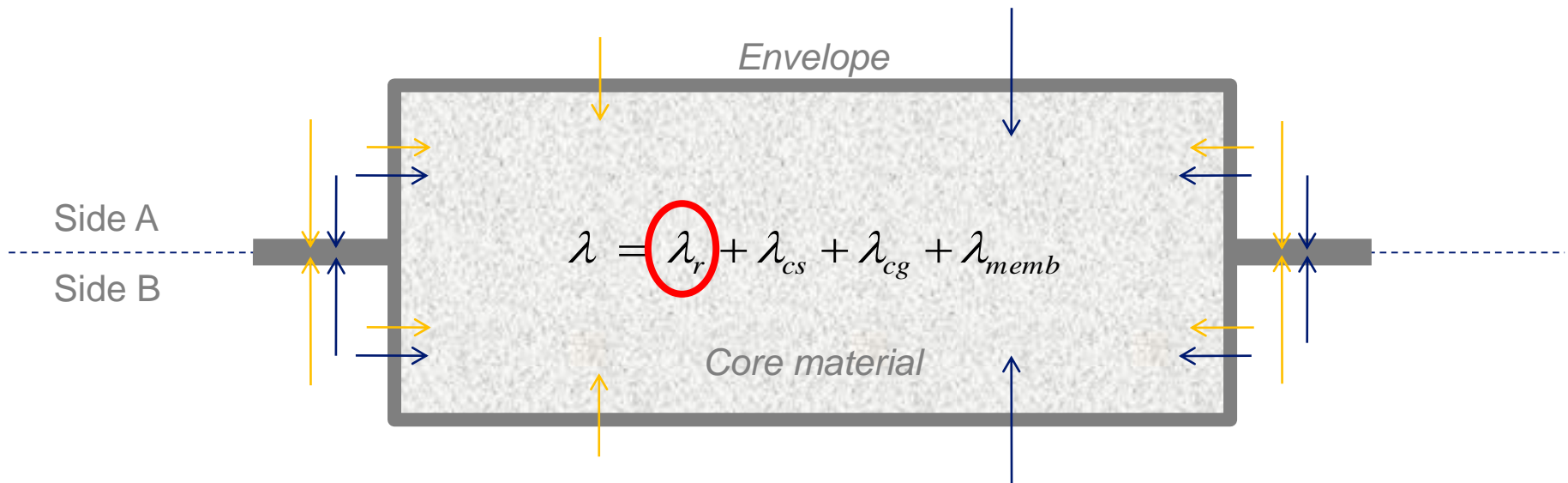
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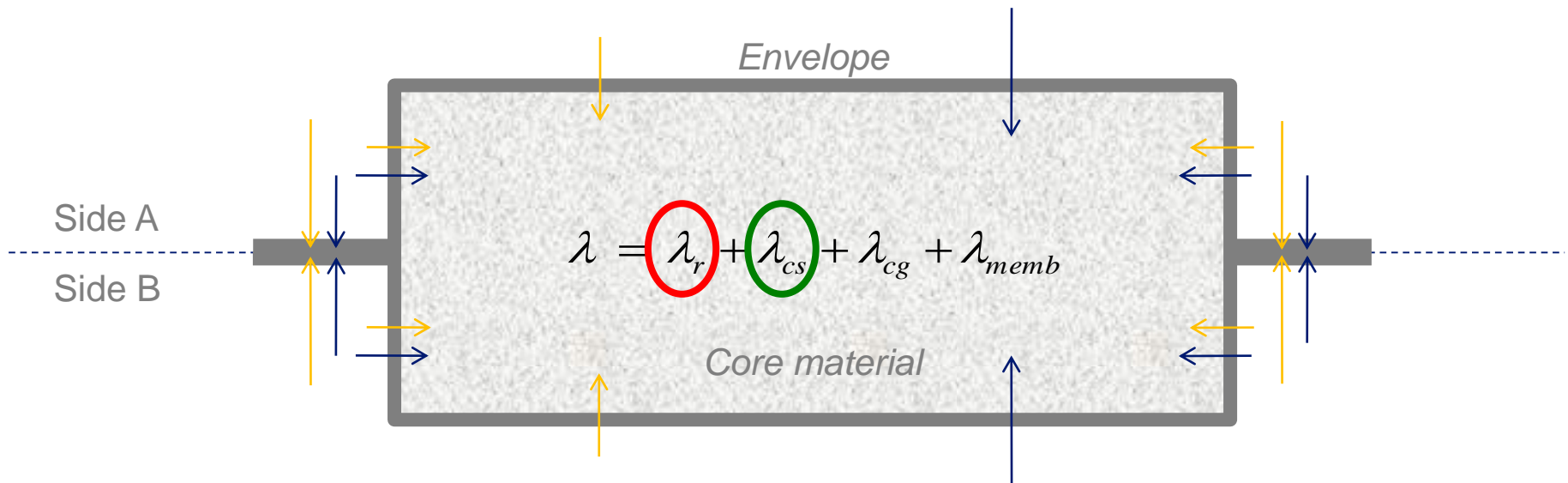
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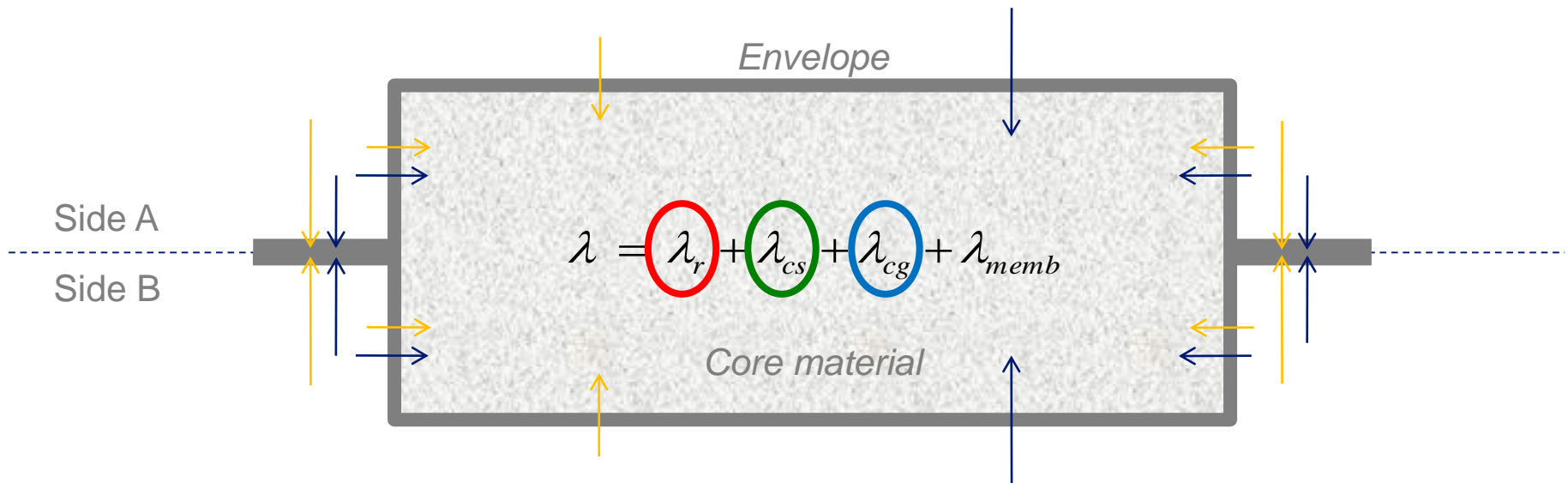
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+ pore size distribution

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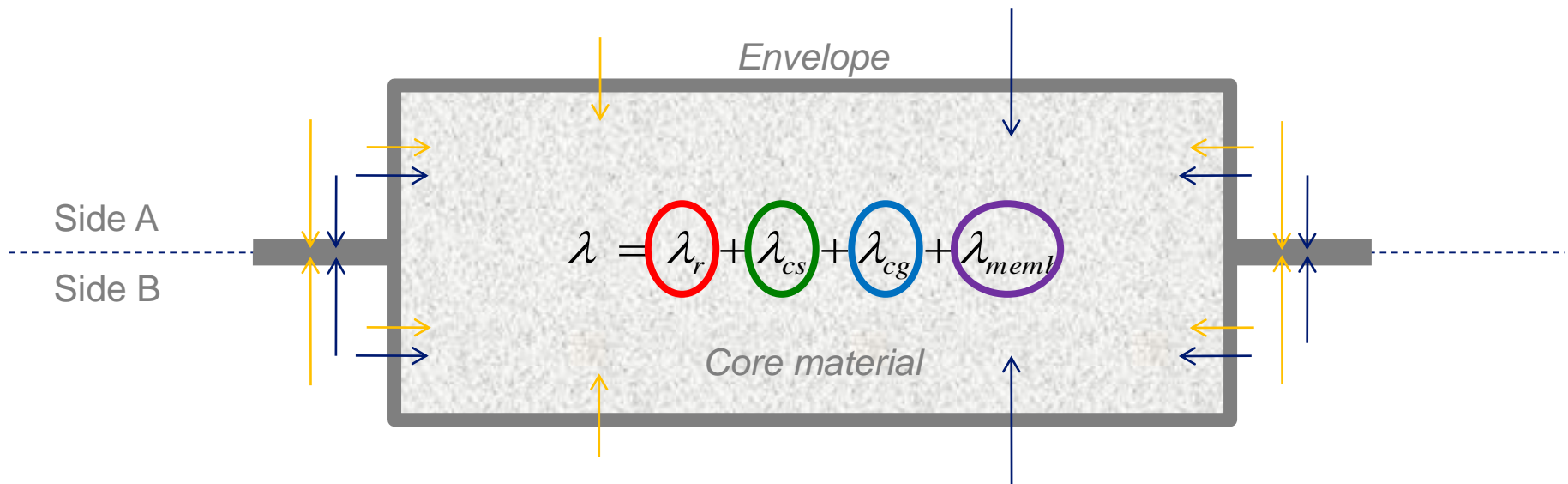
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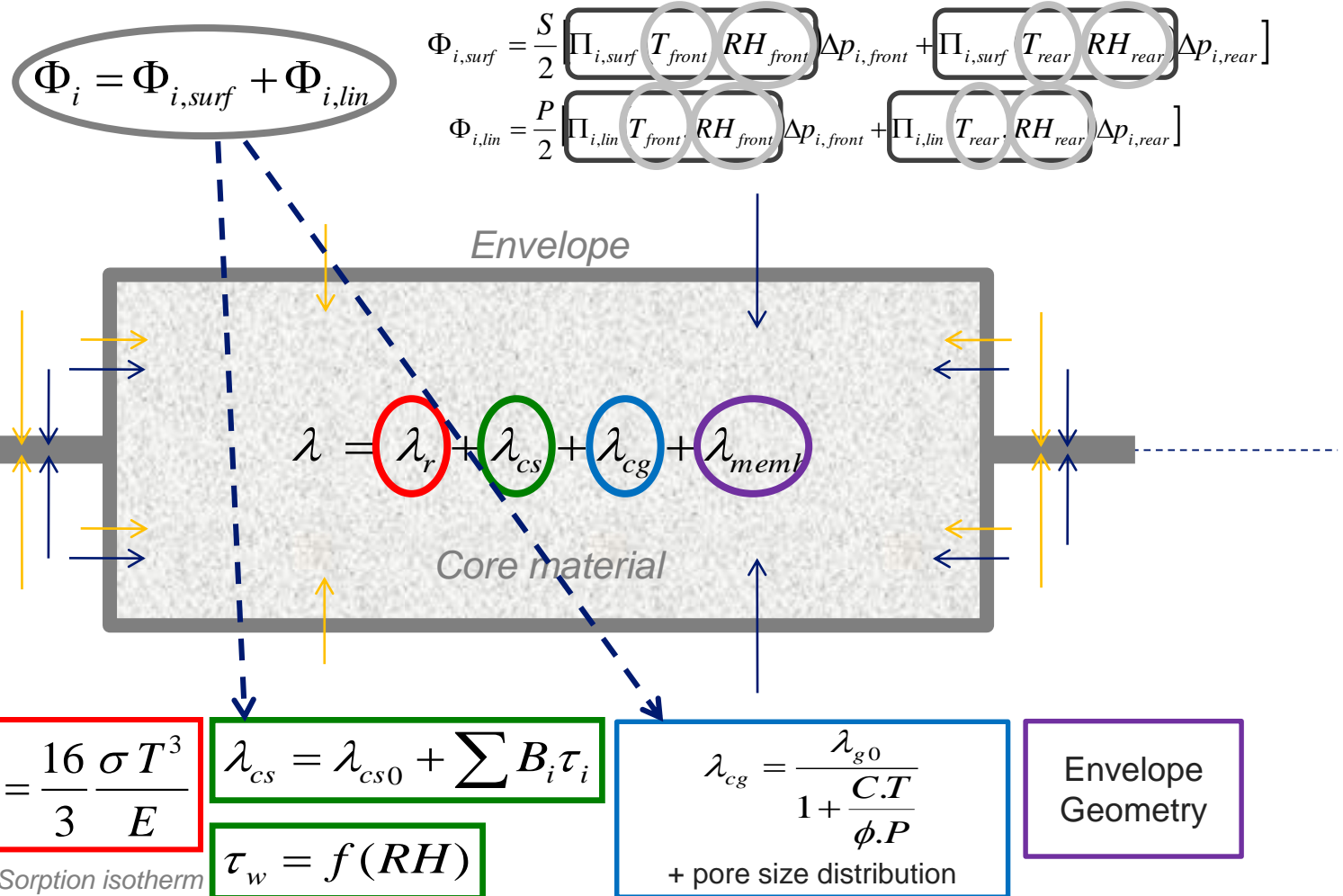
+ pore size distribution

Envelope  
Geometry

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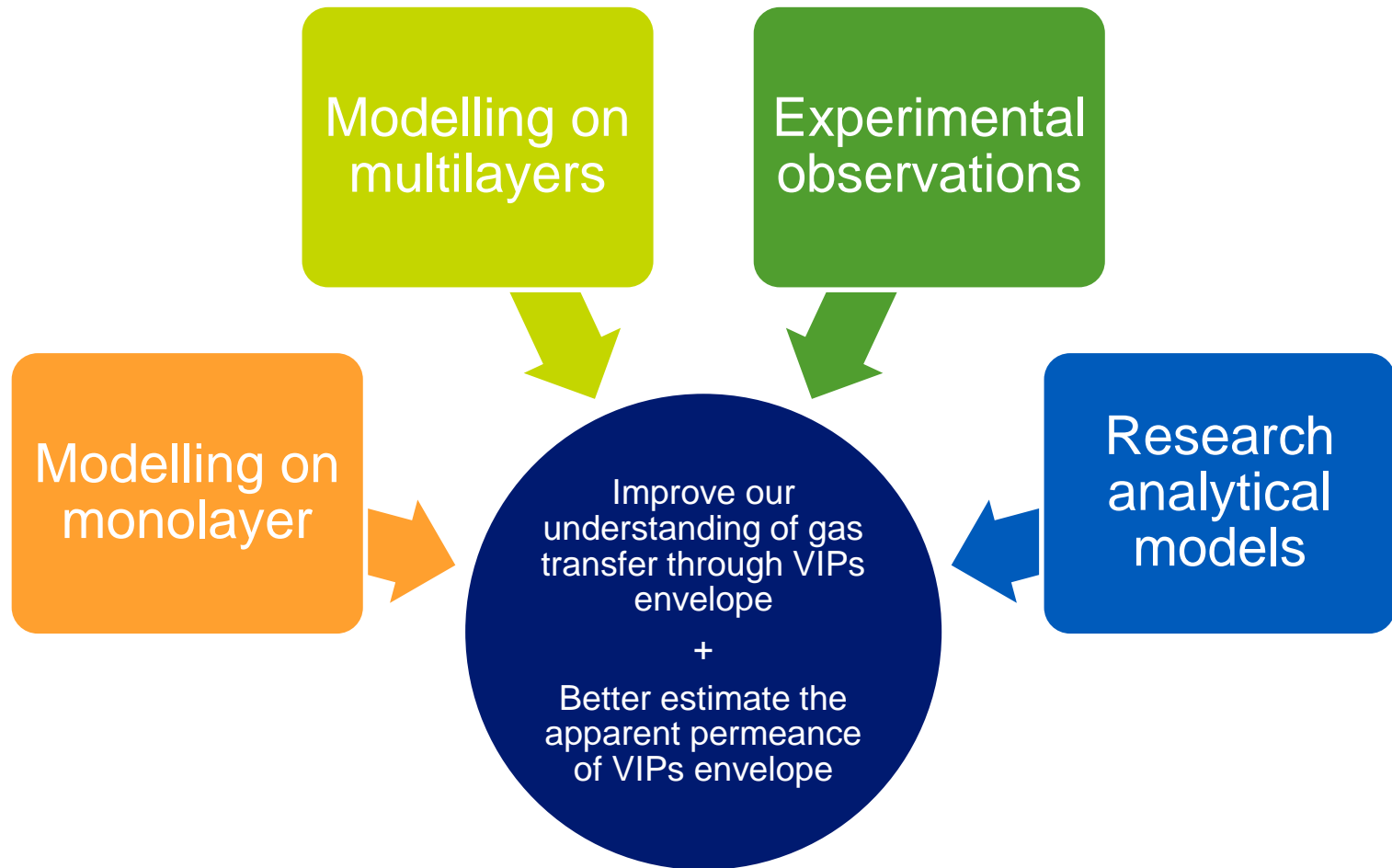




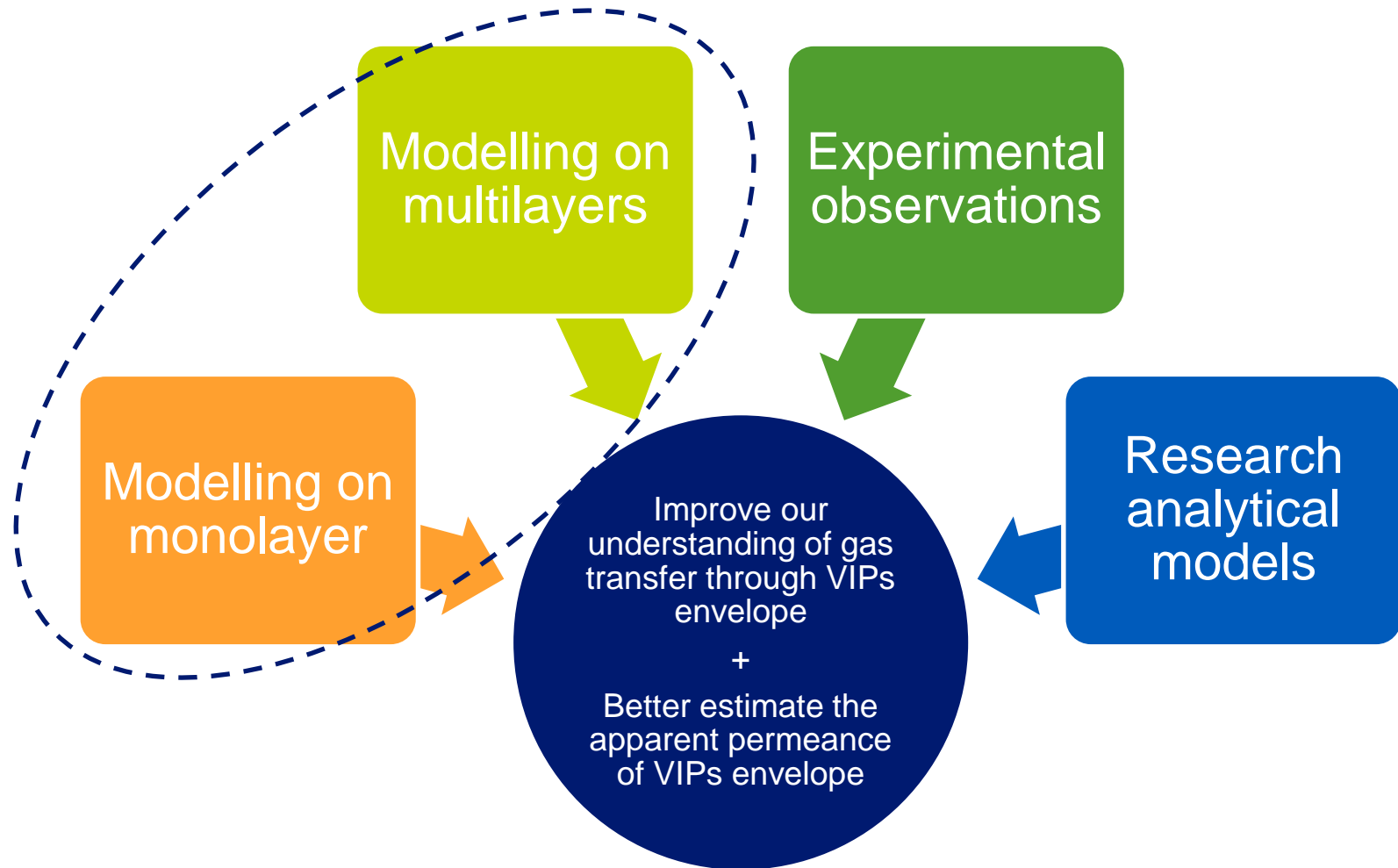
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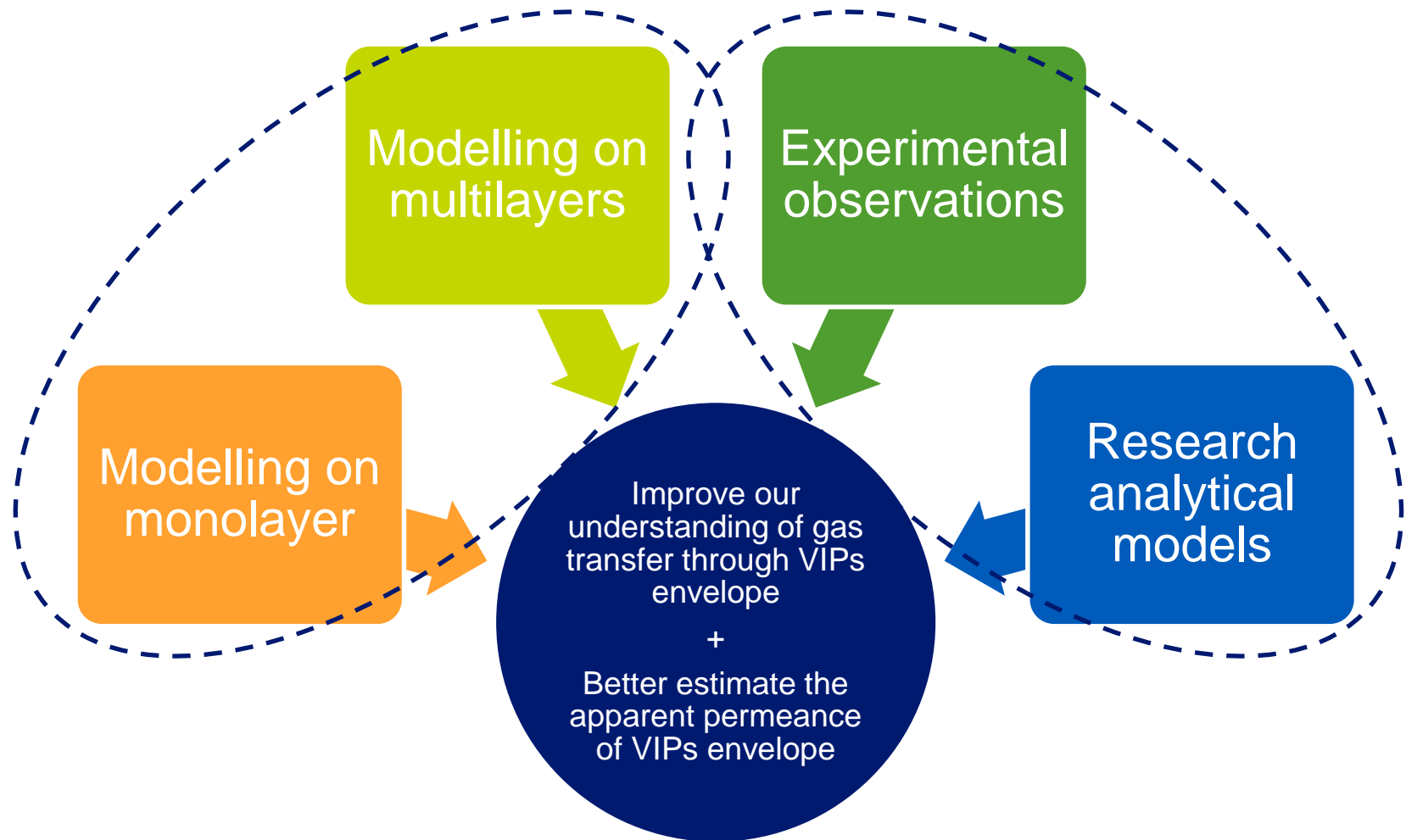
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# MODELLING APPROACH

# BASICS PHENOMENA

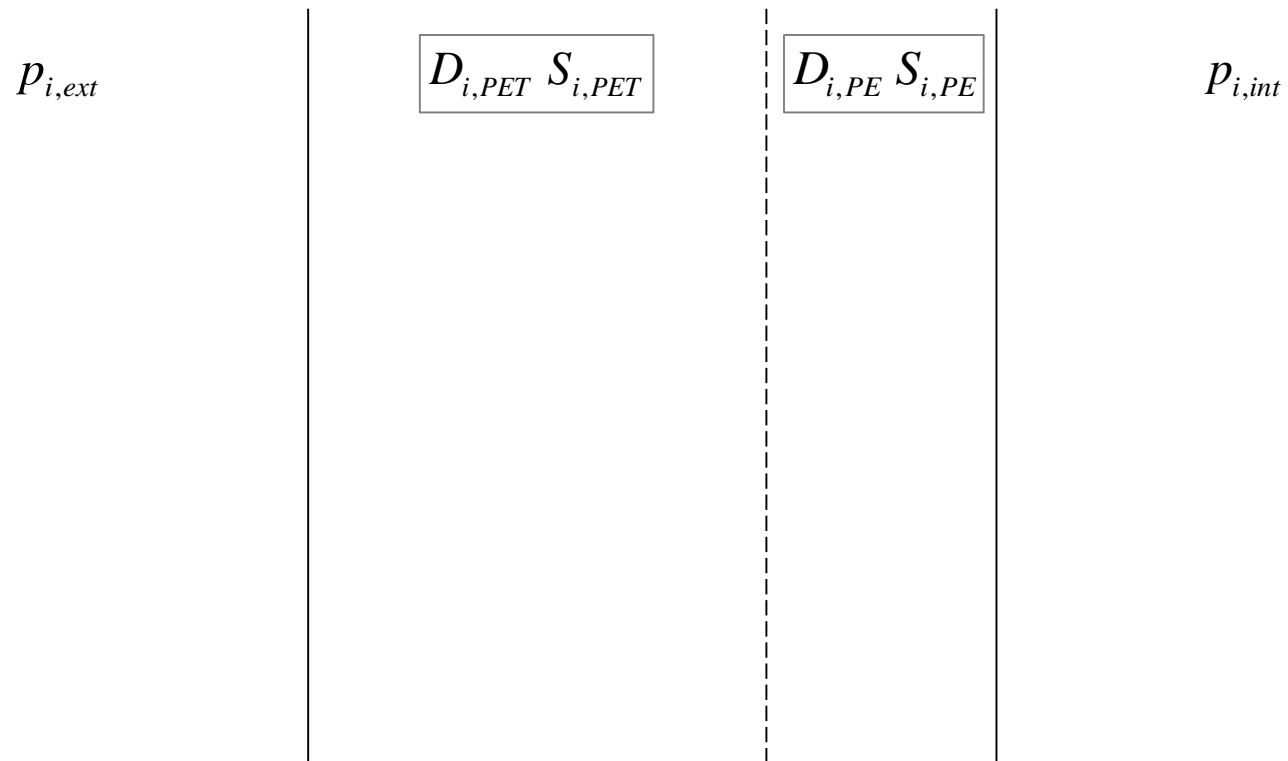
- Dissolution and molecular diffusion

$$D_{i,PET} \ S_{i,PET}$$

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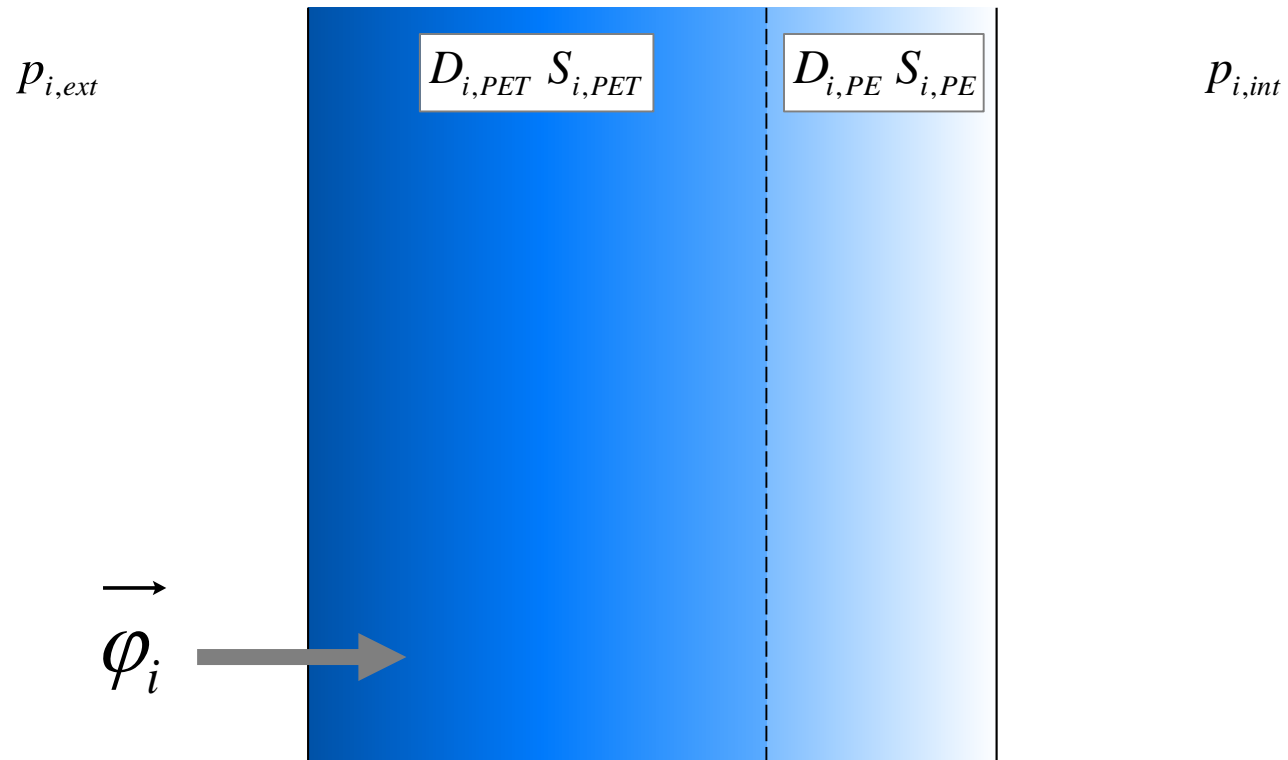
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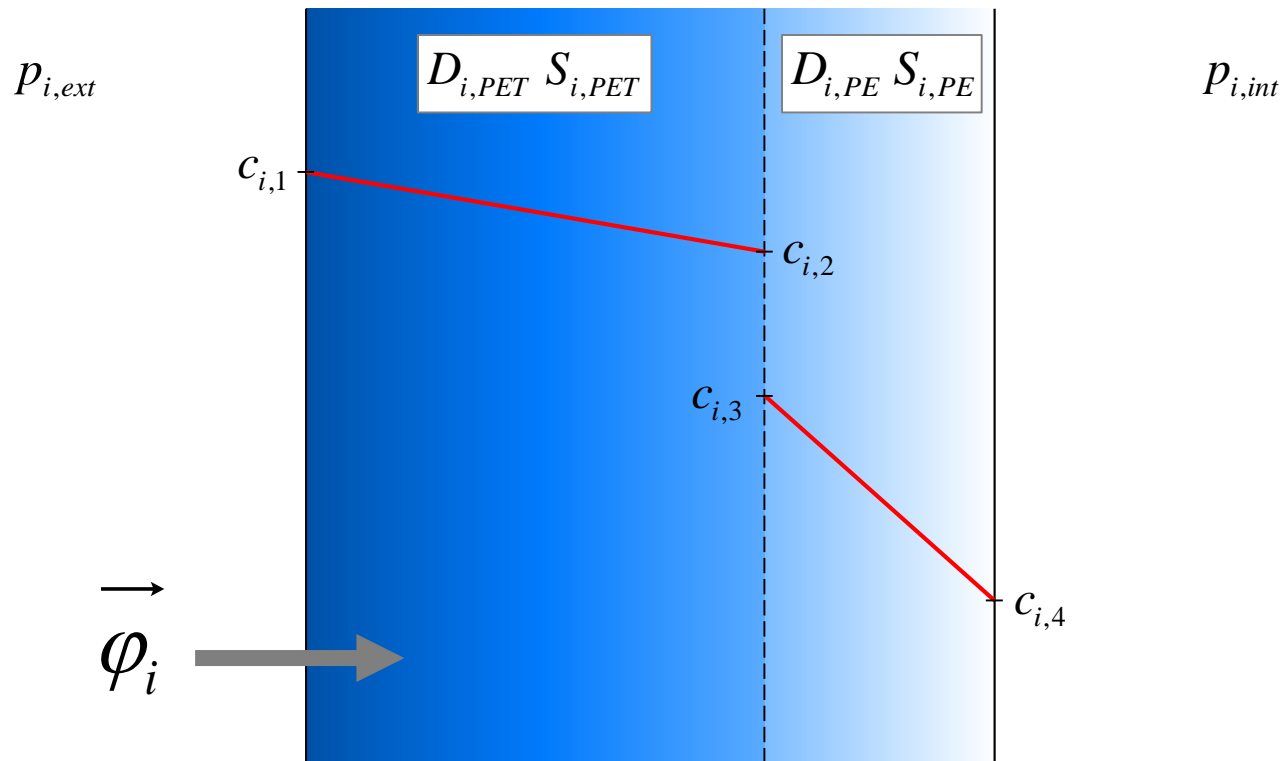
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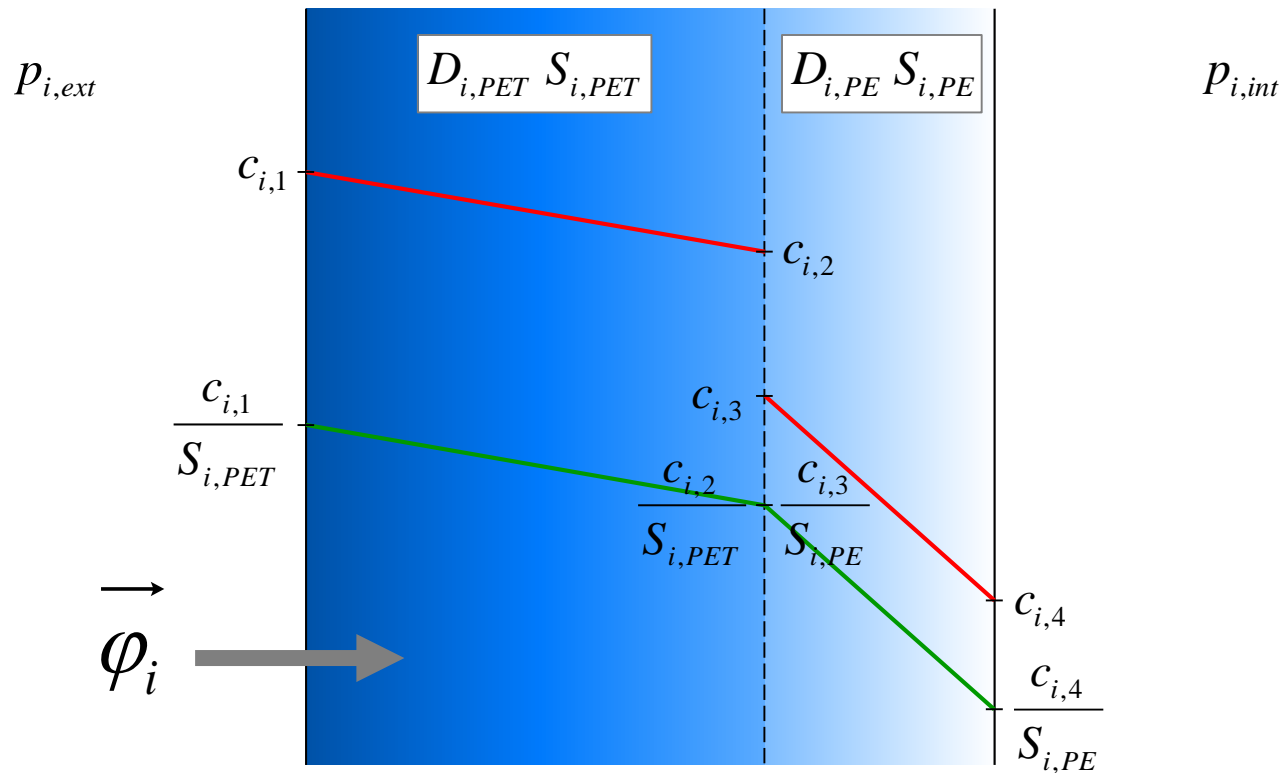
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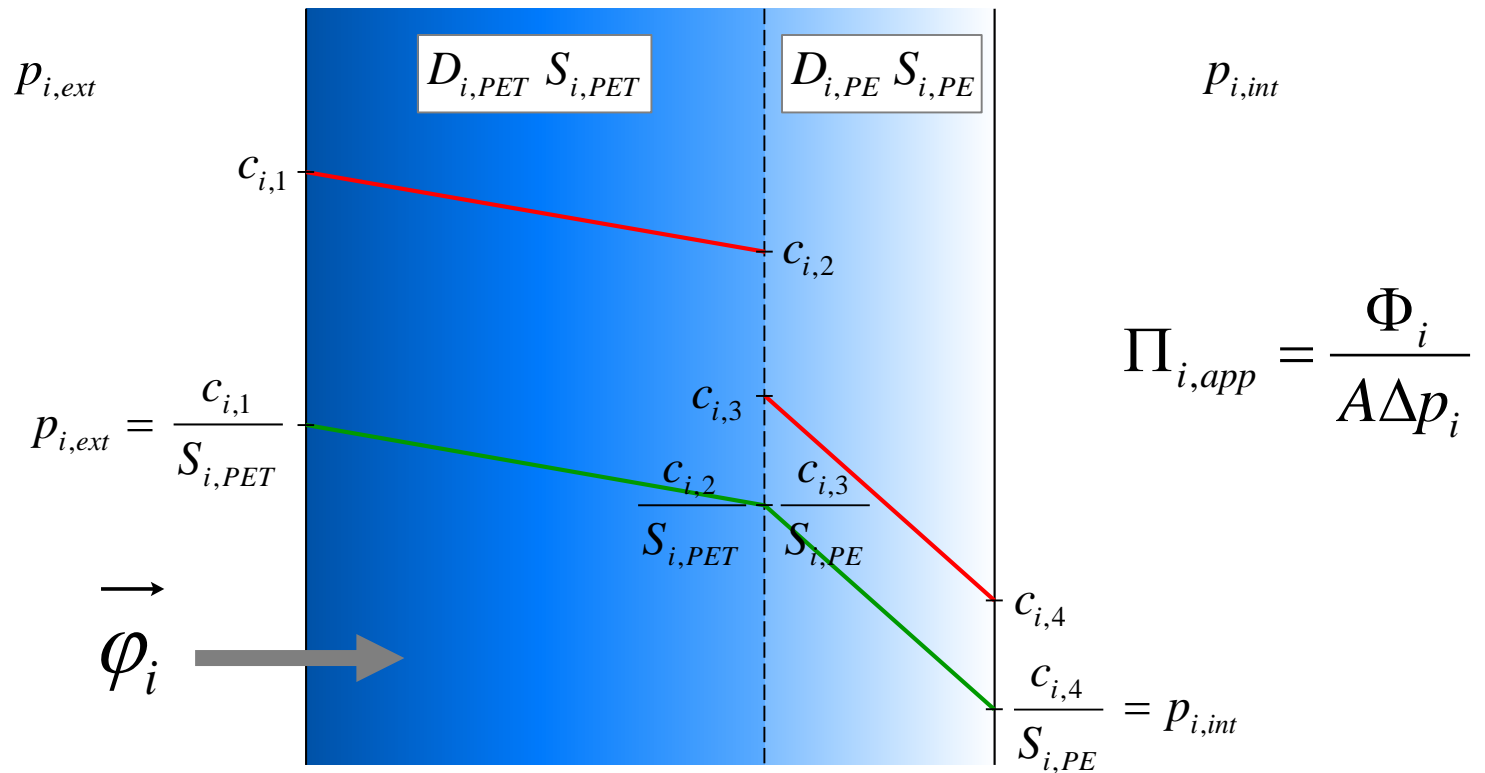
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# BASICS PHENOMENA

- Dissolution and molecular diffusion



# MODELLING APPROACH

- Analogy between mass and heat transfer

Analogy	Heat conduction	Molecular diffusion
Potential	Temperature $T$	Concentration/Solubility $\frac{c_i}{S_{i,j}}$
Flux	$\vec{\phi} = -\lambda \nabla T$ <p>Fourier's law</p>	$\vec{\phi}_i = -D_{i,j} S_{i,j} \nabla \frac{c_i}{S_{i,j}}$ <p>Fick's law</p>
Transitional regime	$\frac{\partial T}{\partial t} = \frac{\lambda}{\rho c_p} \nabla^2 T$ <p>Energy conservation</p>	$\frac{\partial}{\partial t} \left( \frac{c_i}{S_{i,j}} \right) = D_{i,j} \nabla^2 \frac{c_i}{S_{i,j}}$ <p>Mass conservation</p>

$$\lambda \leftrightarrow D_{i,j} S_{i,j}$$

$$\rho c_p \leftrightarrow S_{i,j}$$

# MAIN STEPS OF MODELLING

## ■ MESHING

- Simail®
- SMESH module from SALOME® Platform



## ■ THERMAL CALCULATION

- SYRTHES® software developed by EDF



## ■ PARAMETRIC STUDIES

- YACS module from SALOME® Platform



## ■ POST-PROCESSING

- EnSight® post-processing software



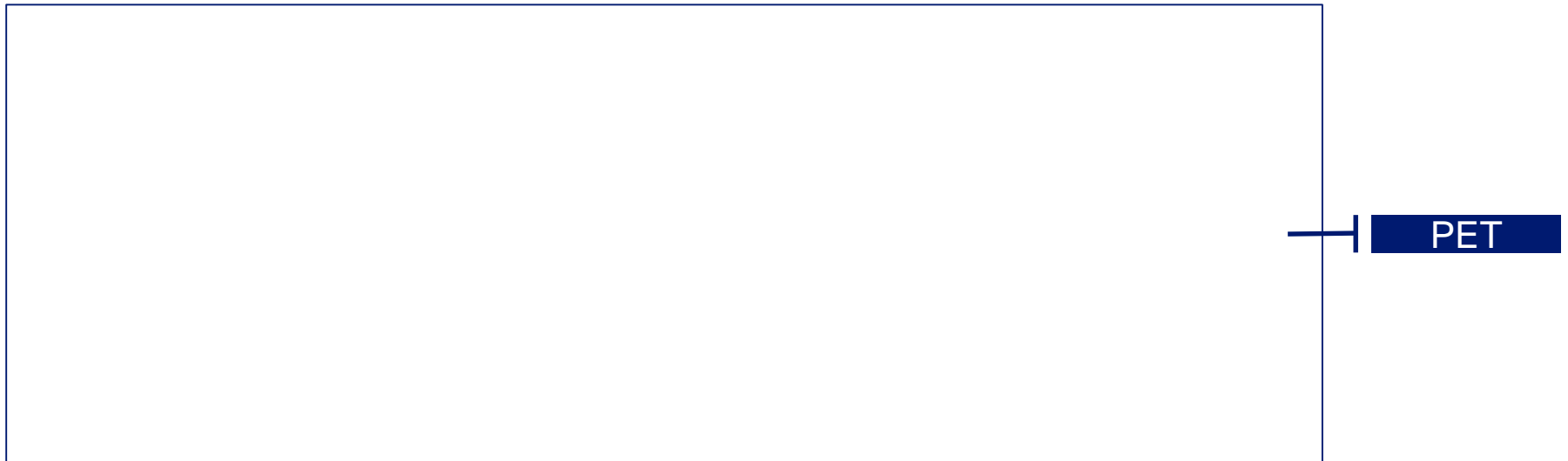


# SOME RESULTS WITH PET M1F

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- PET M1F 80nm

$p_{i,ext}$



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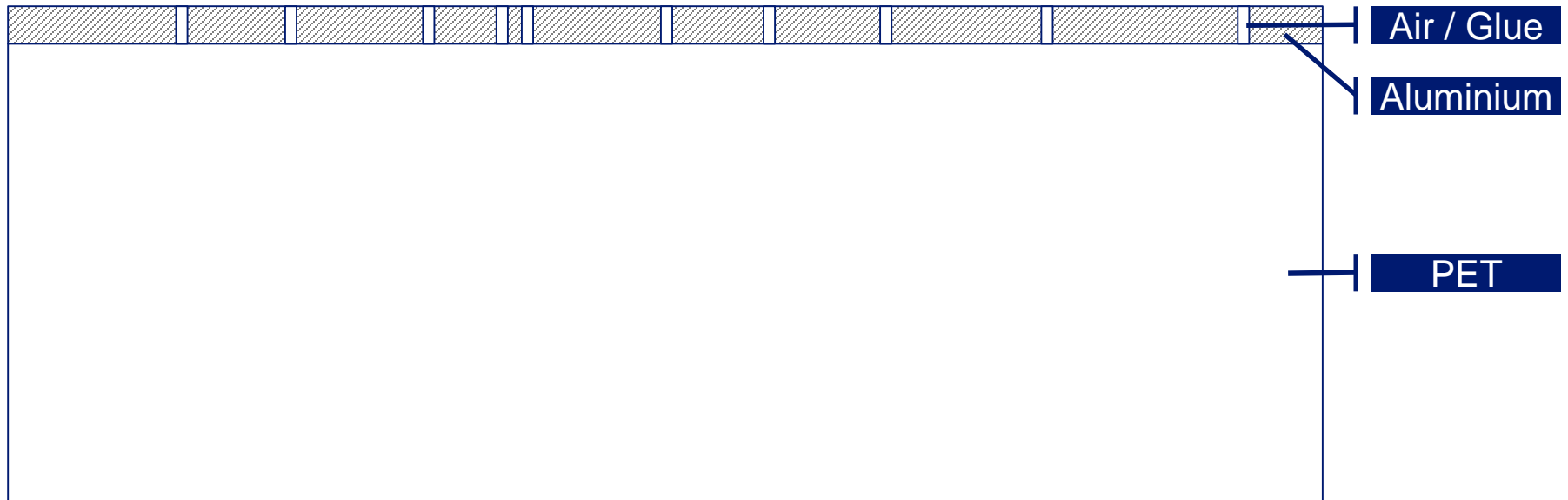
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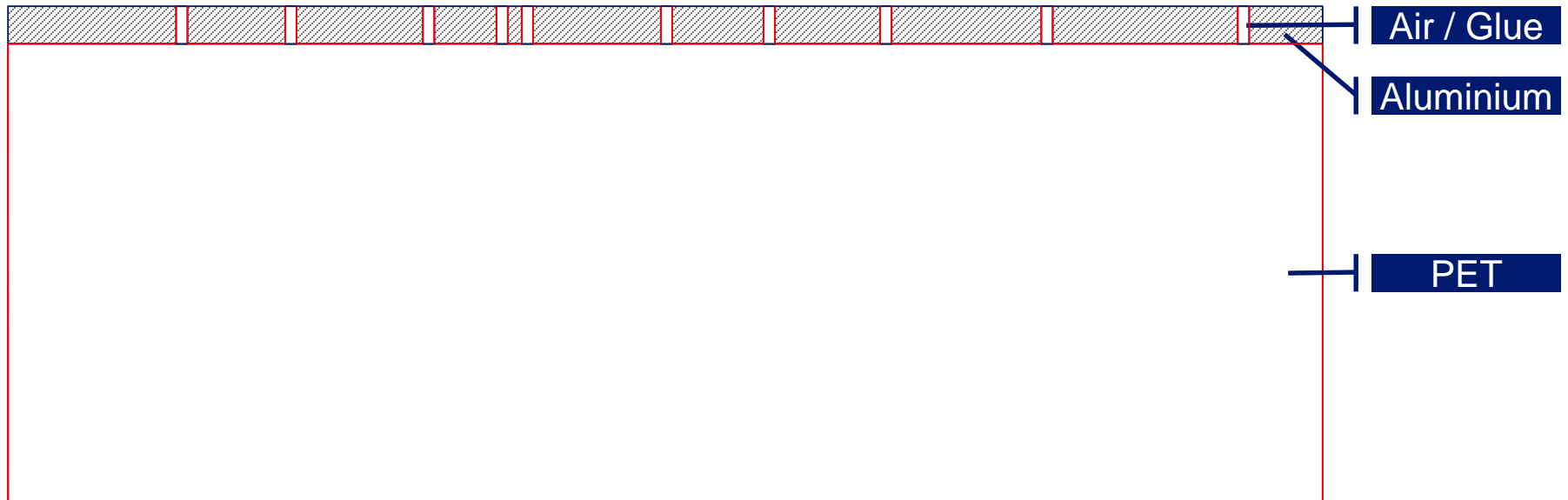


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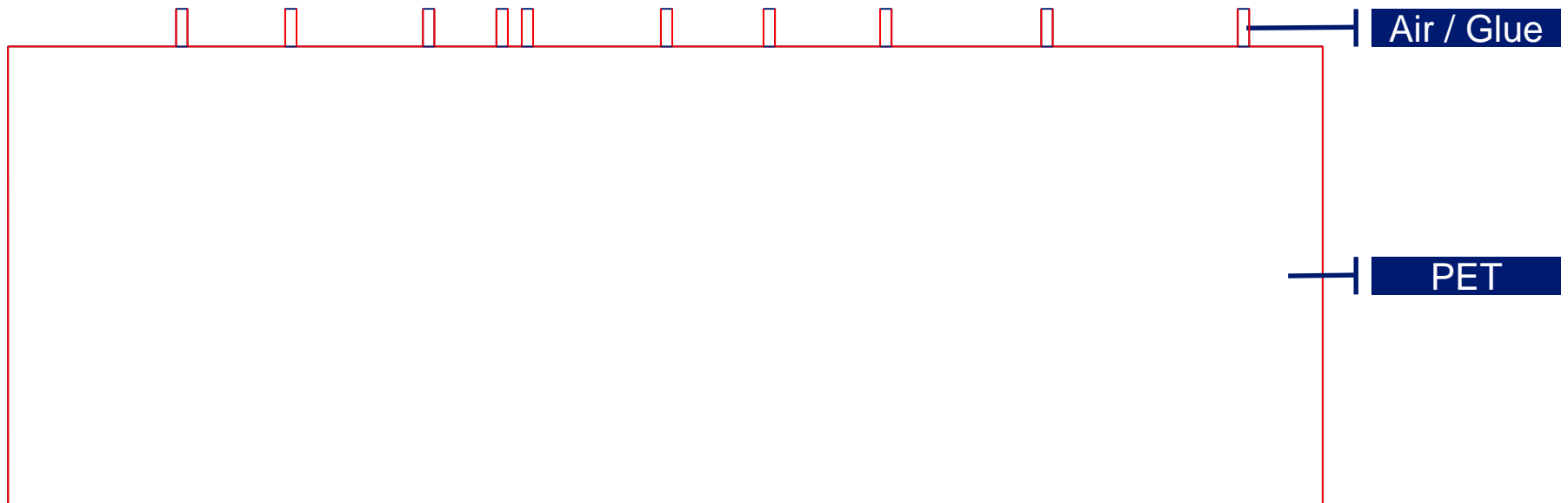
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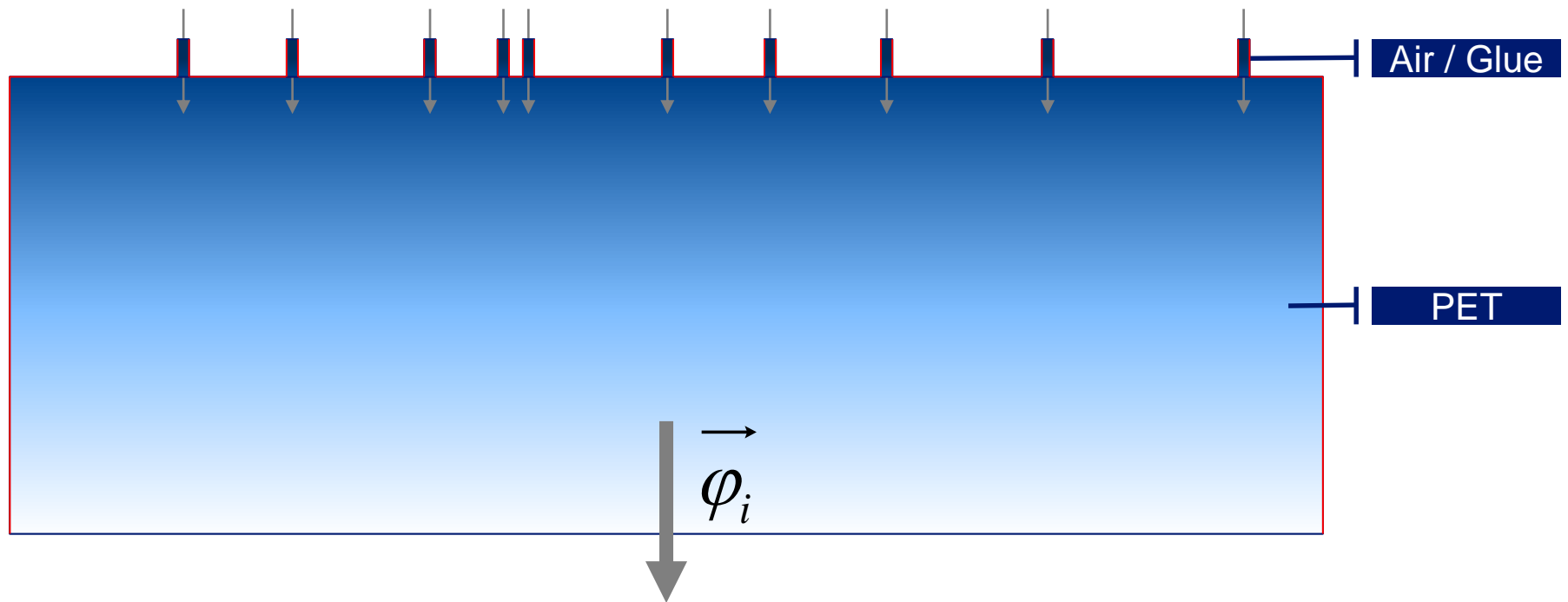
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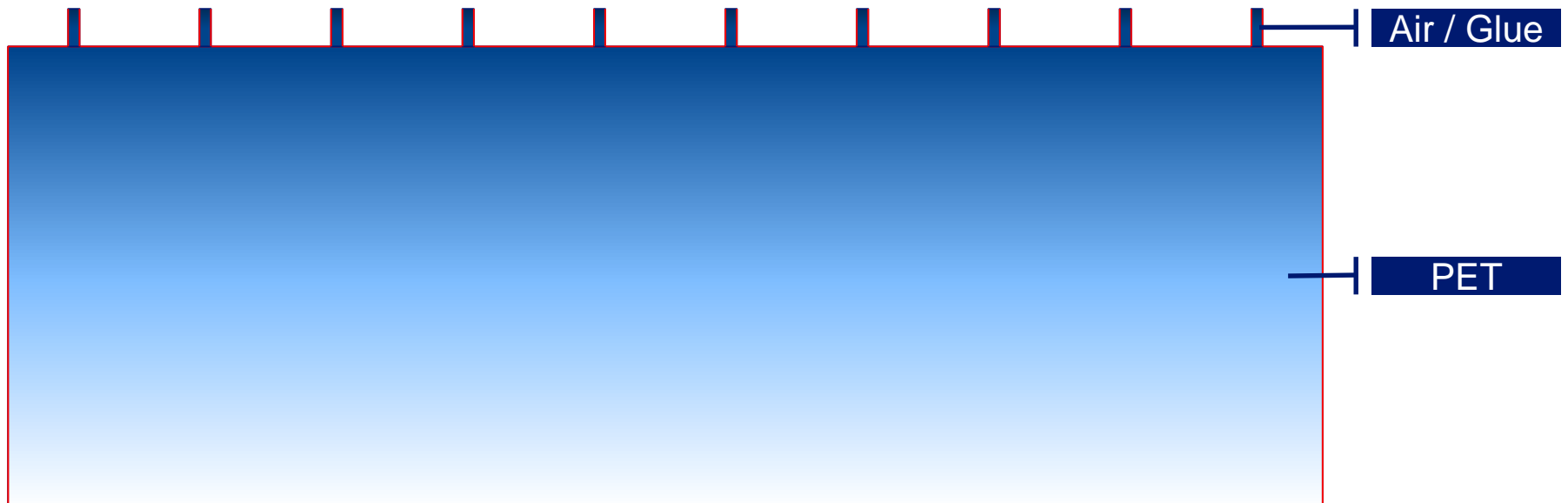
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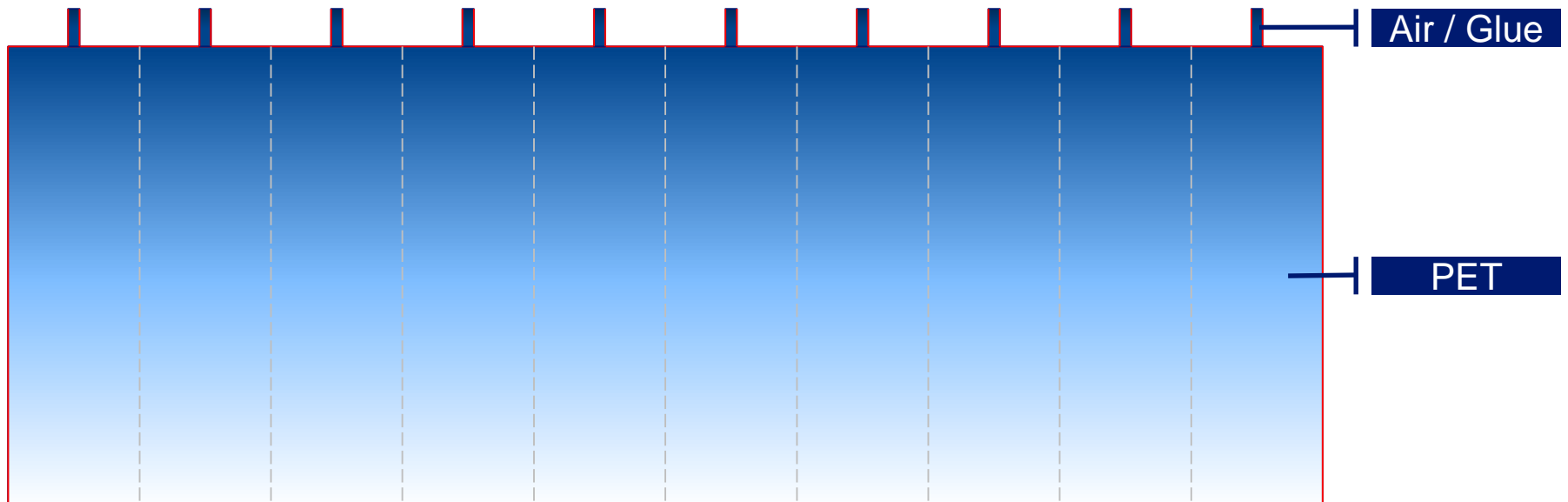
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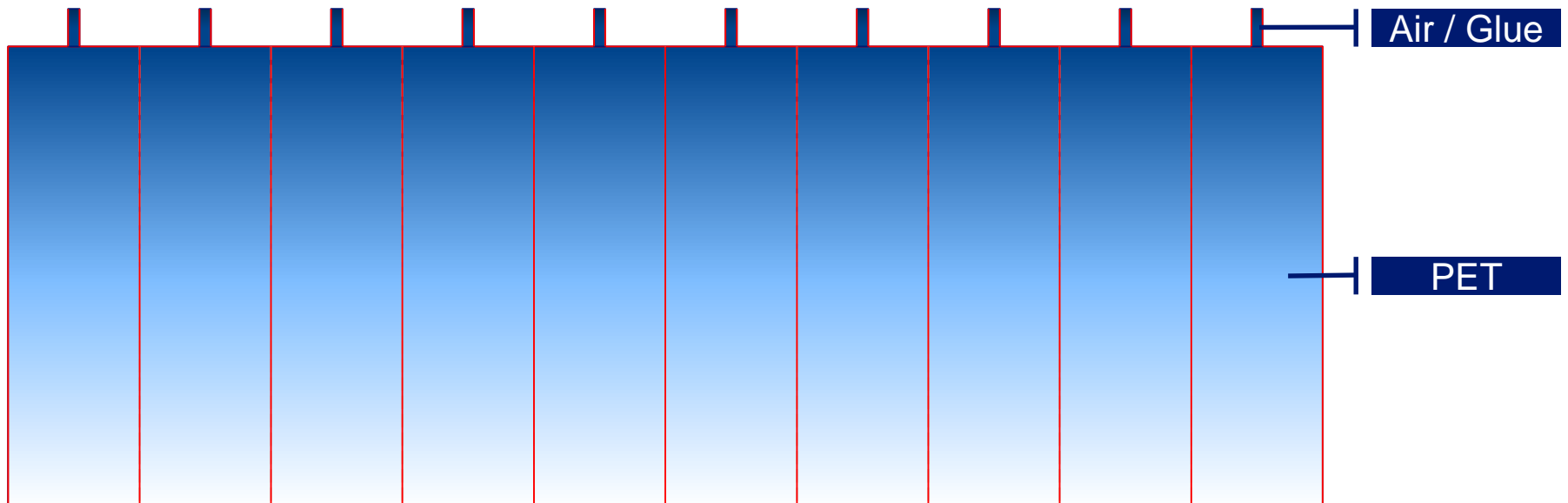
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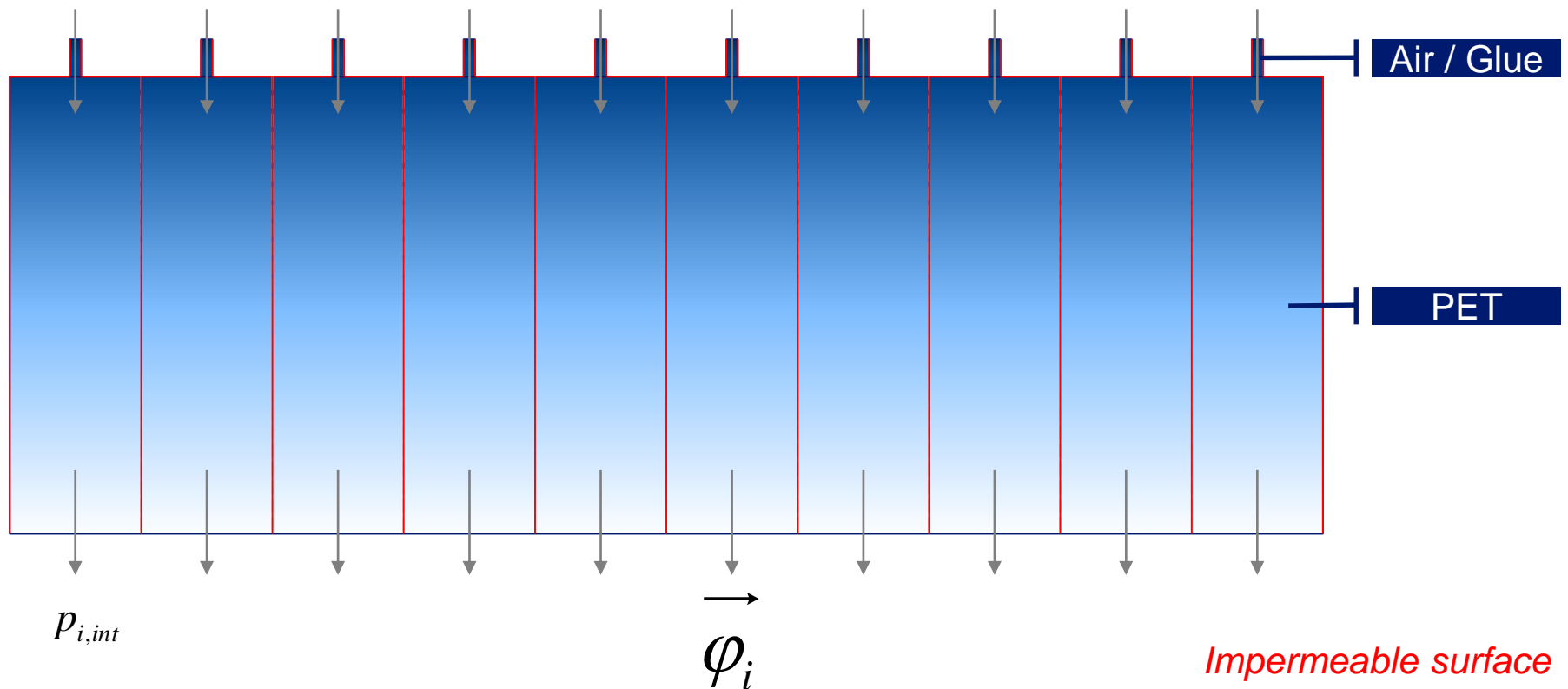
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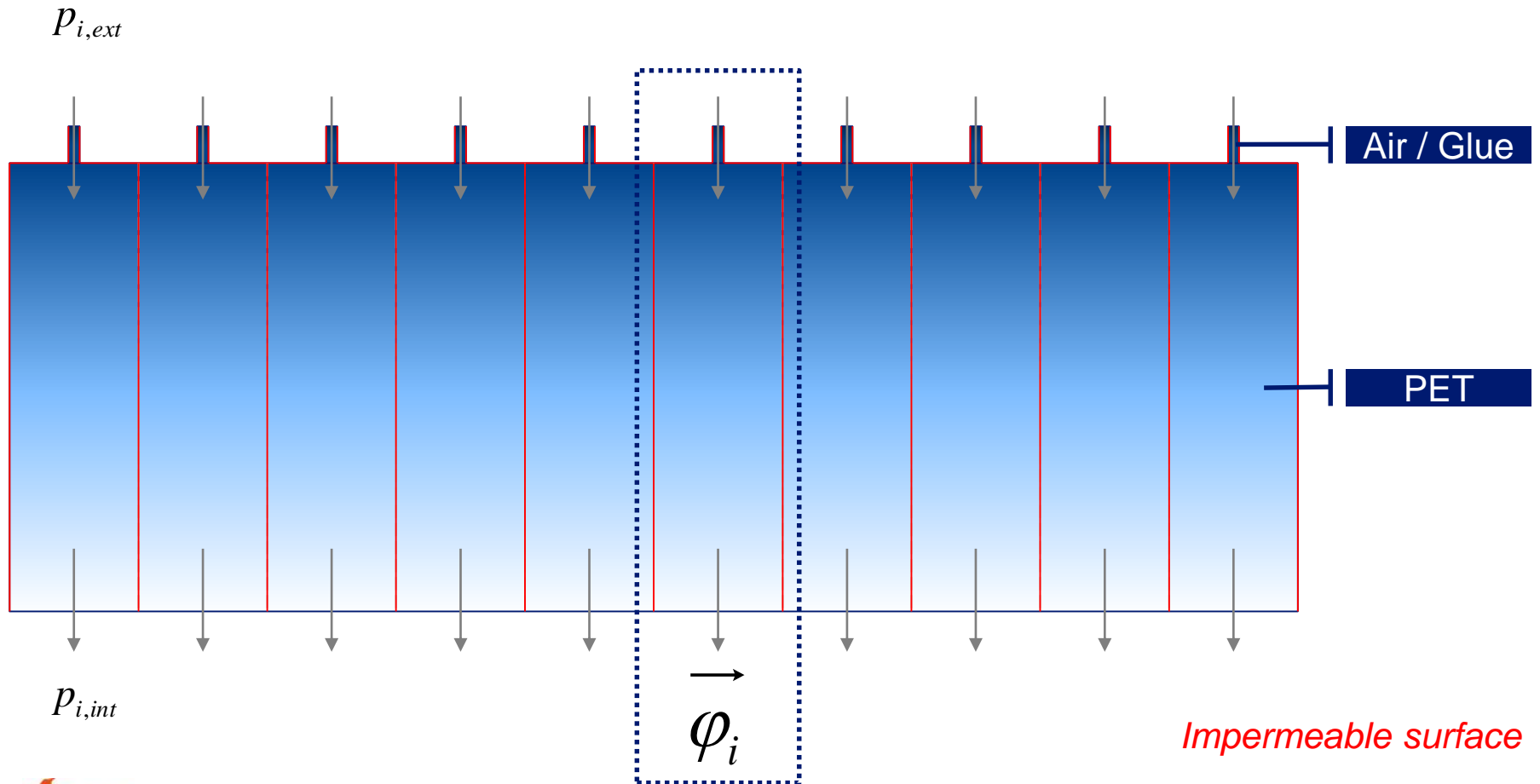
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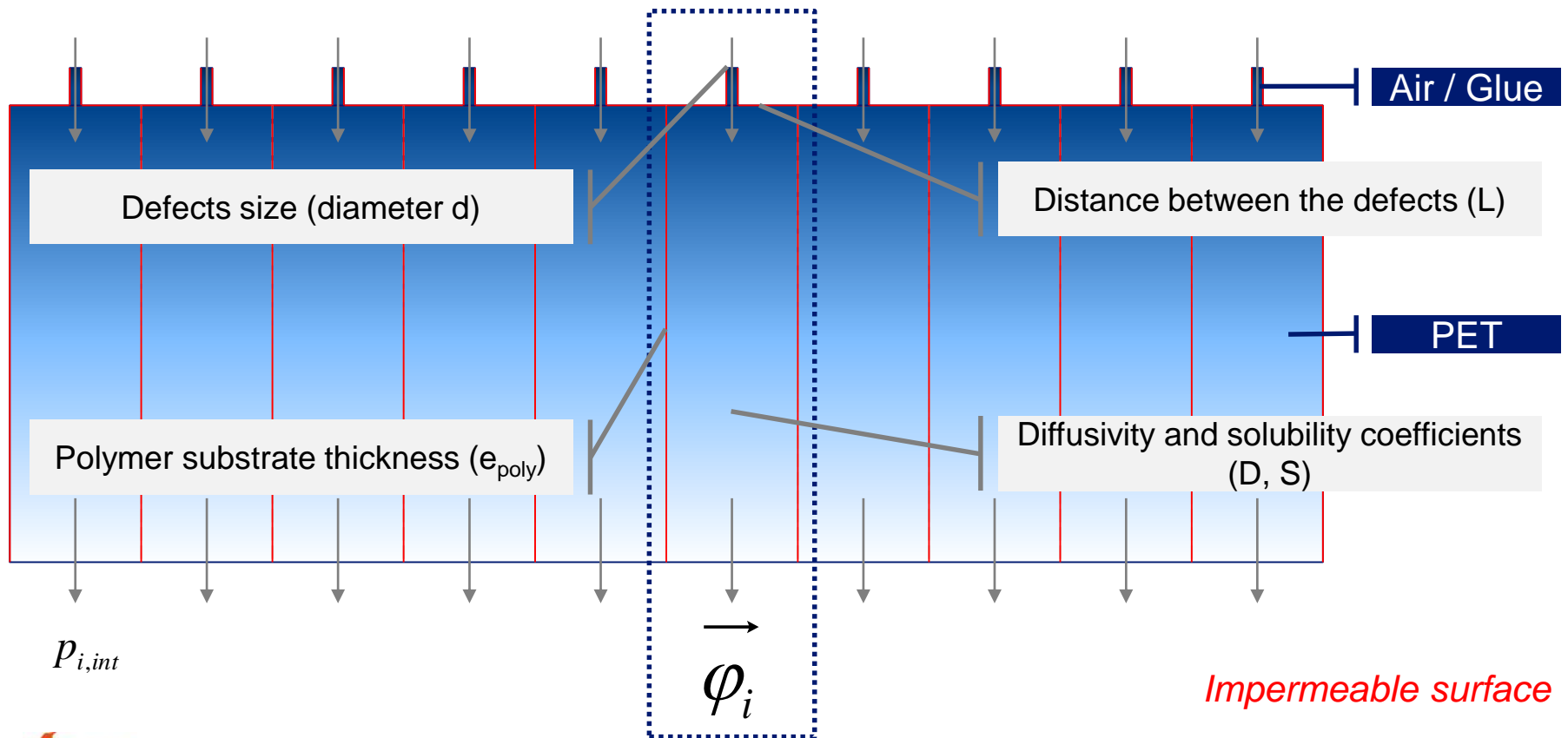
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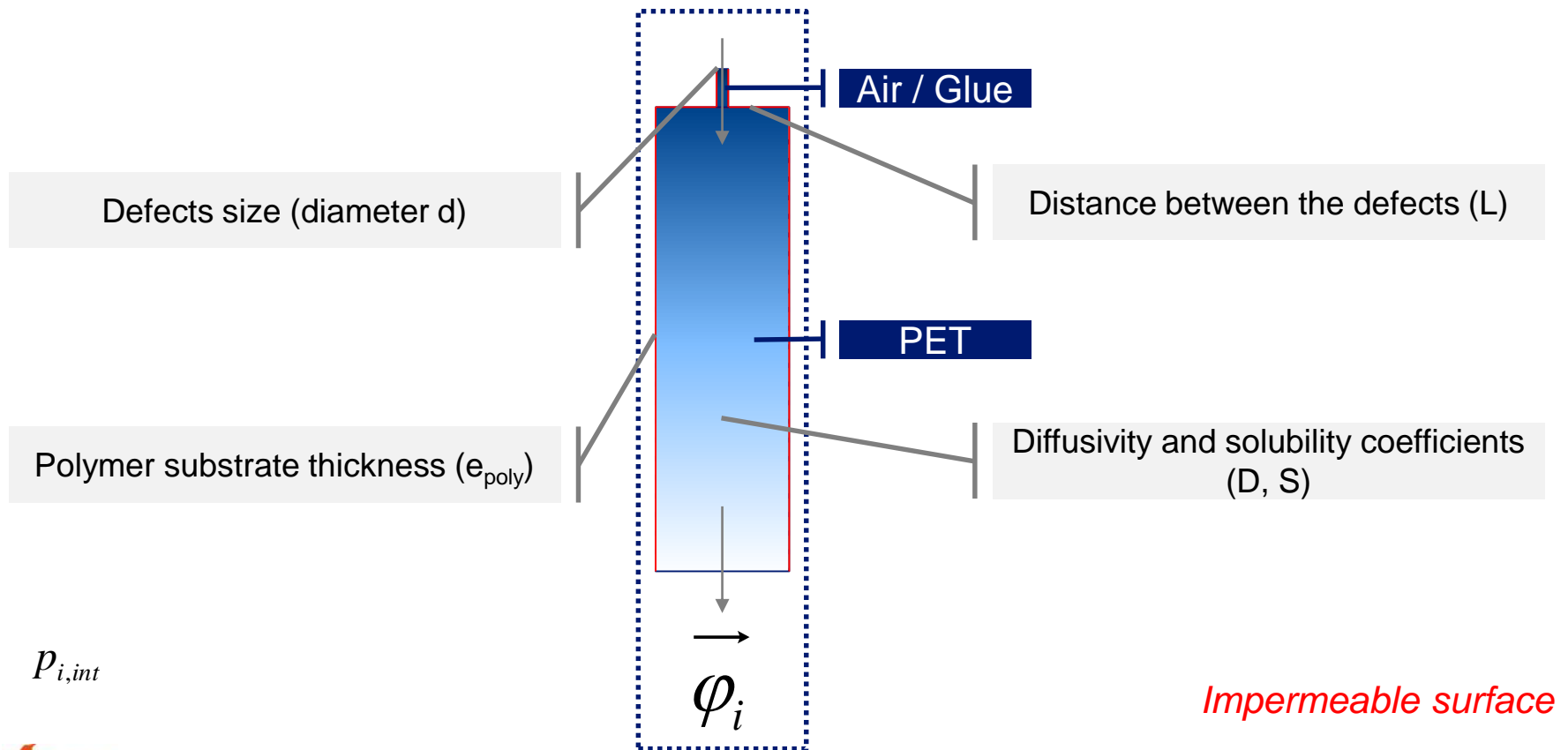
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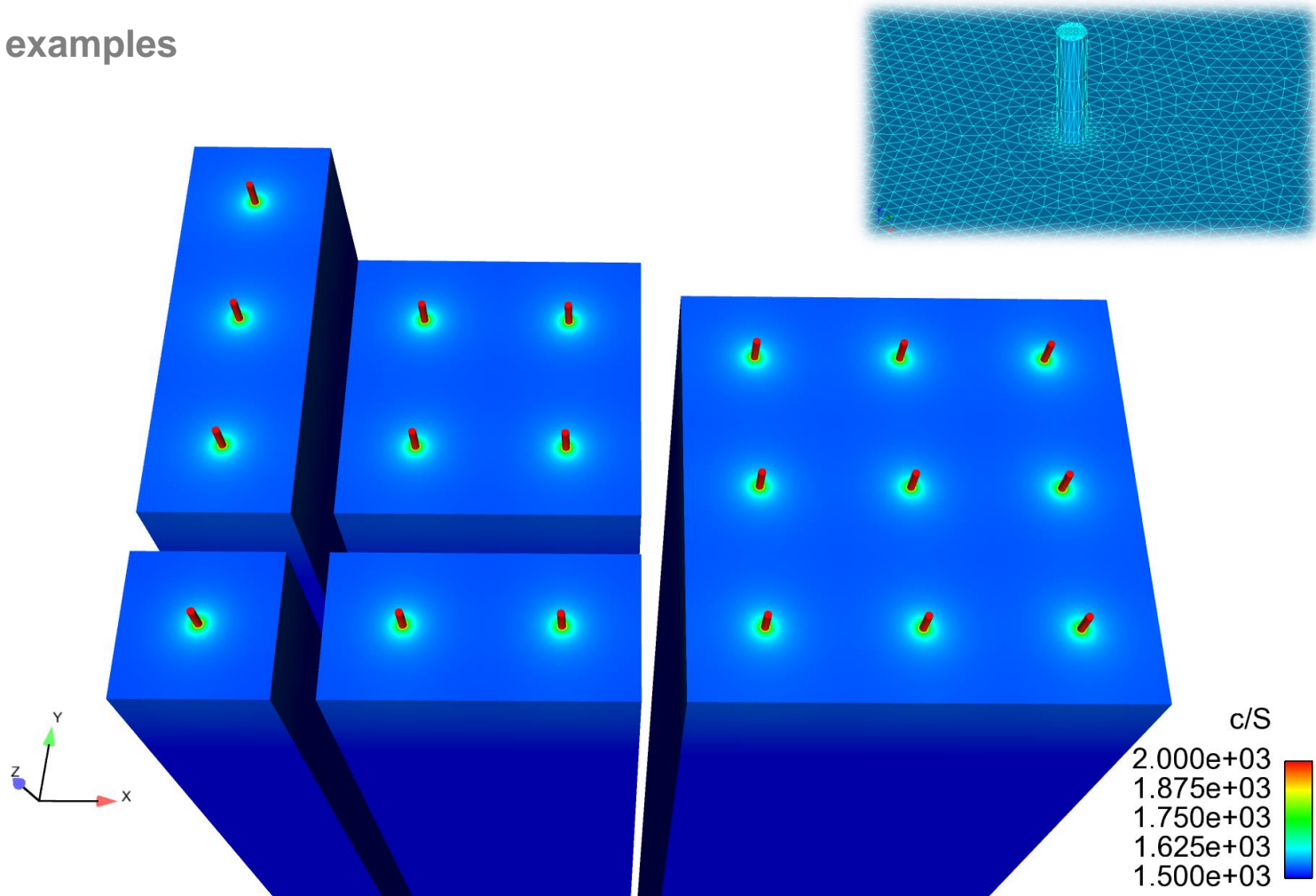
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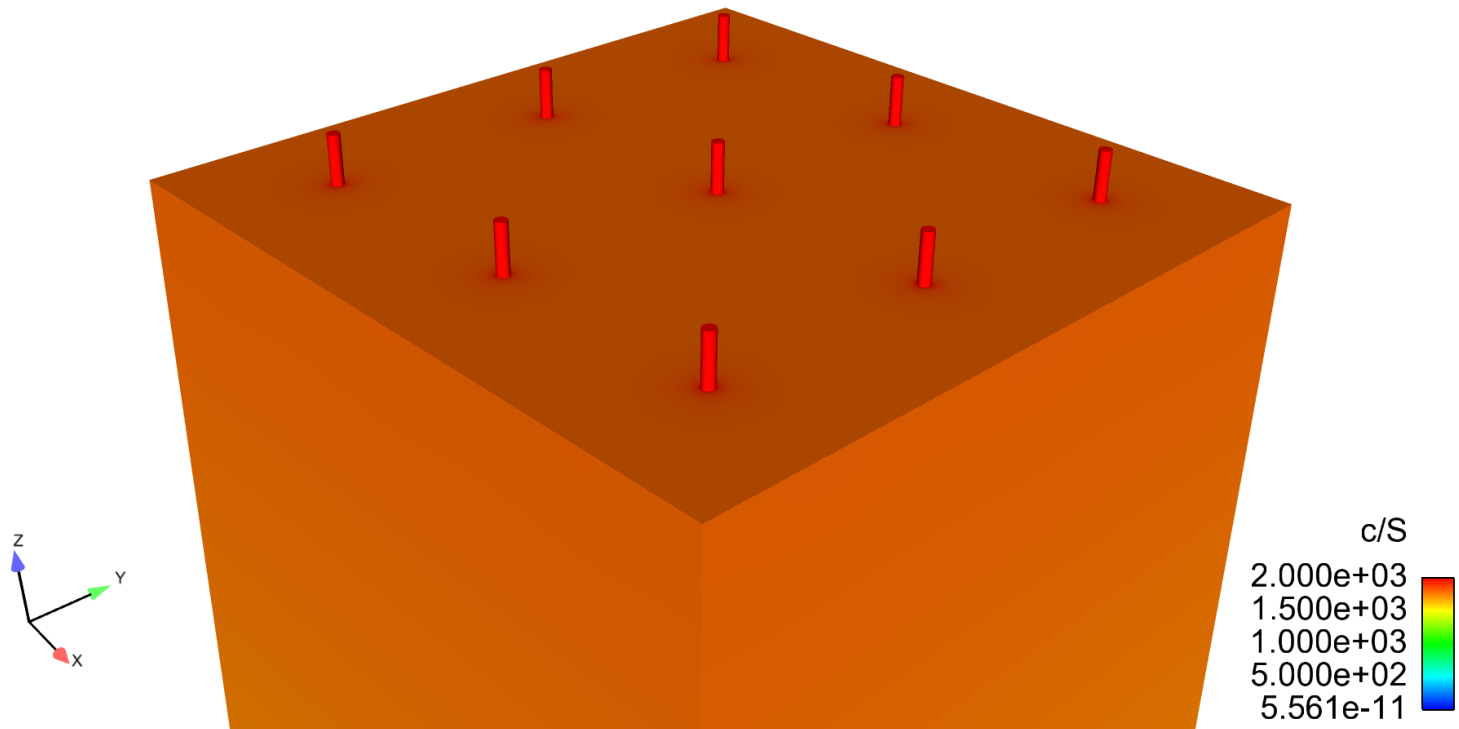
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- Mesh examples



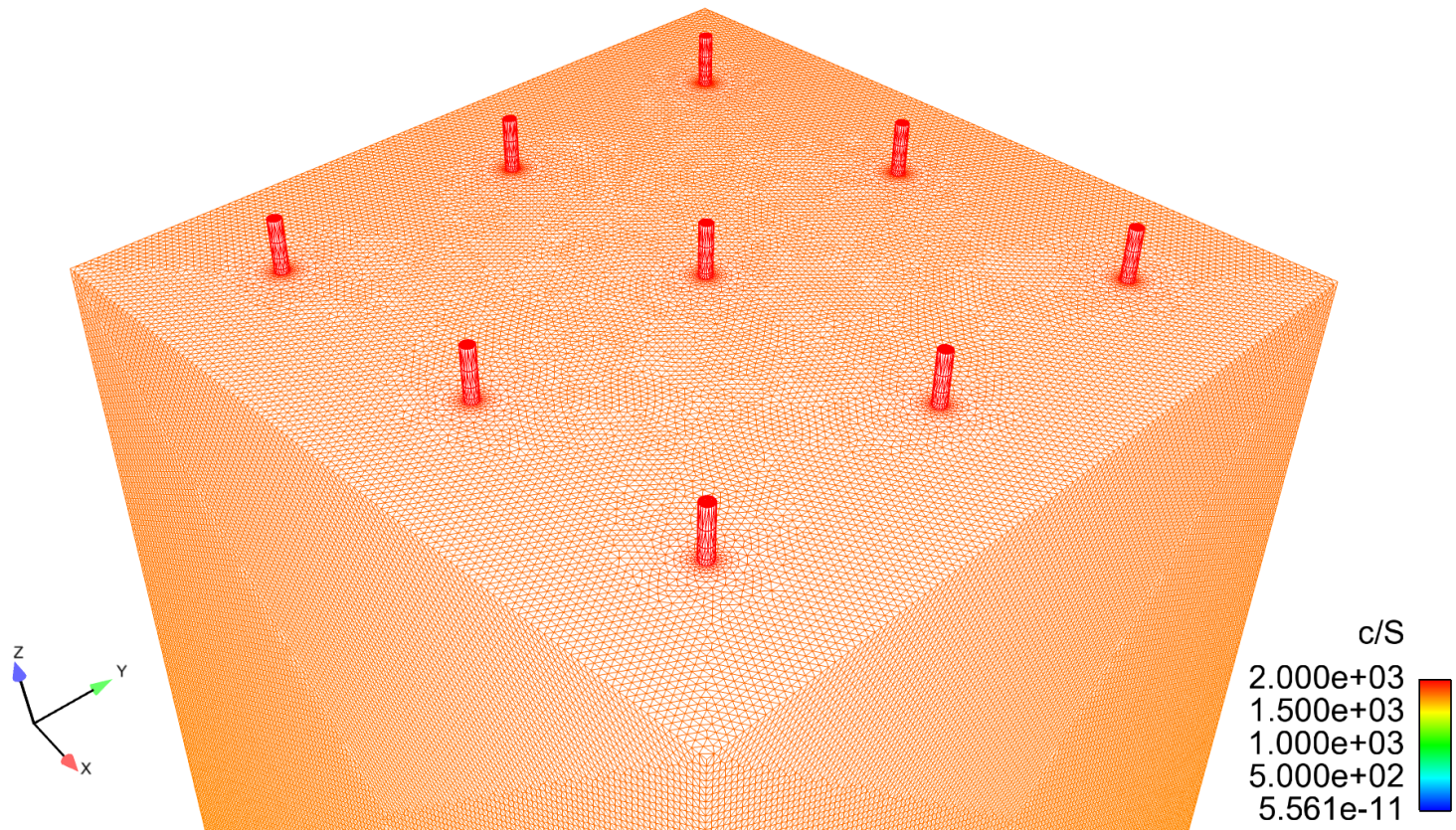
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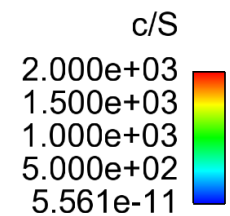
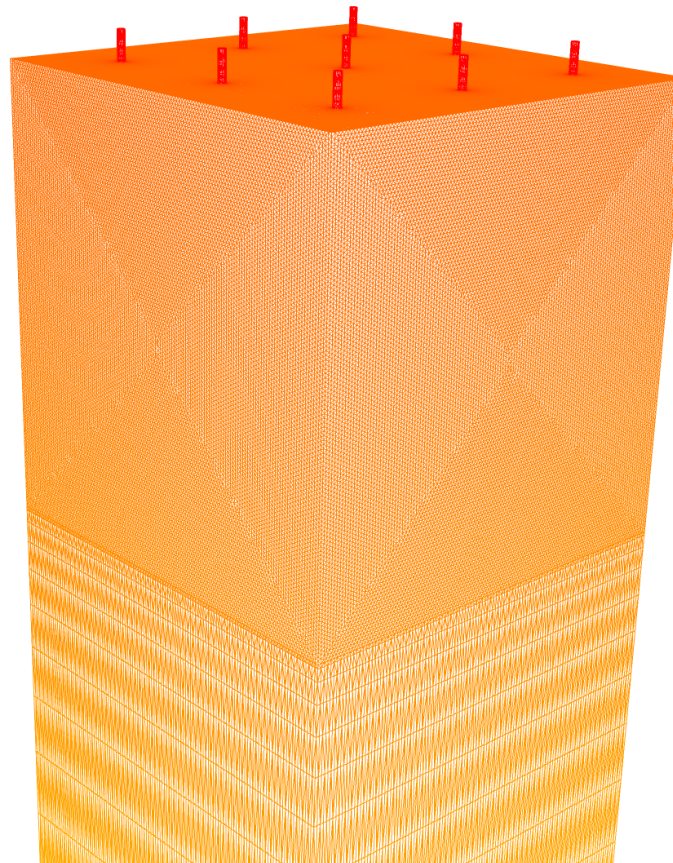
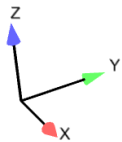
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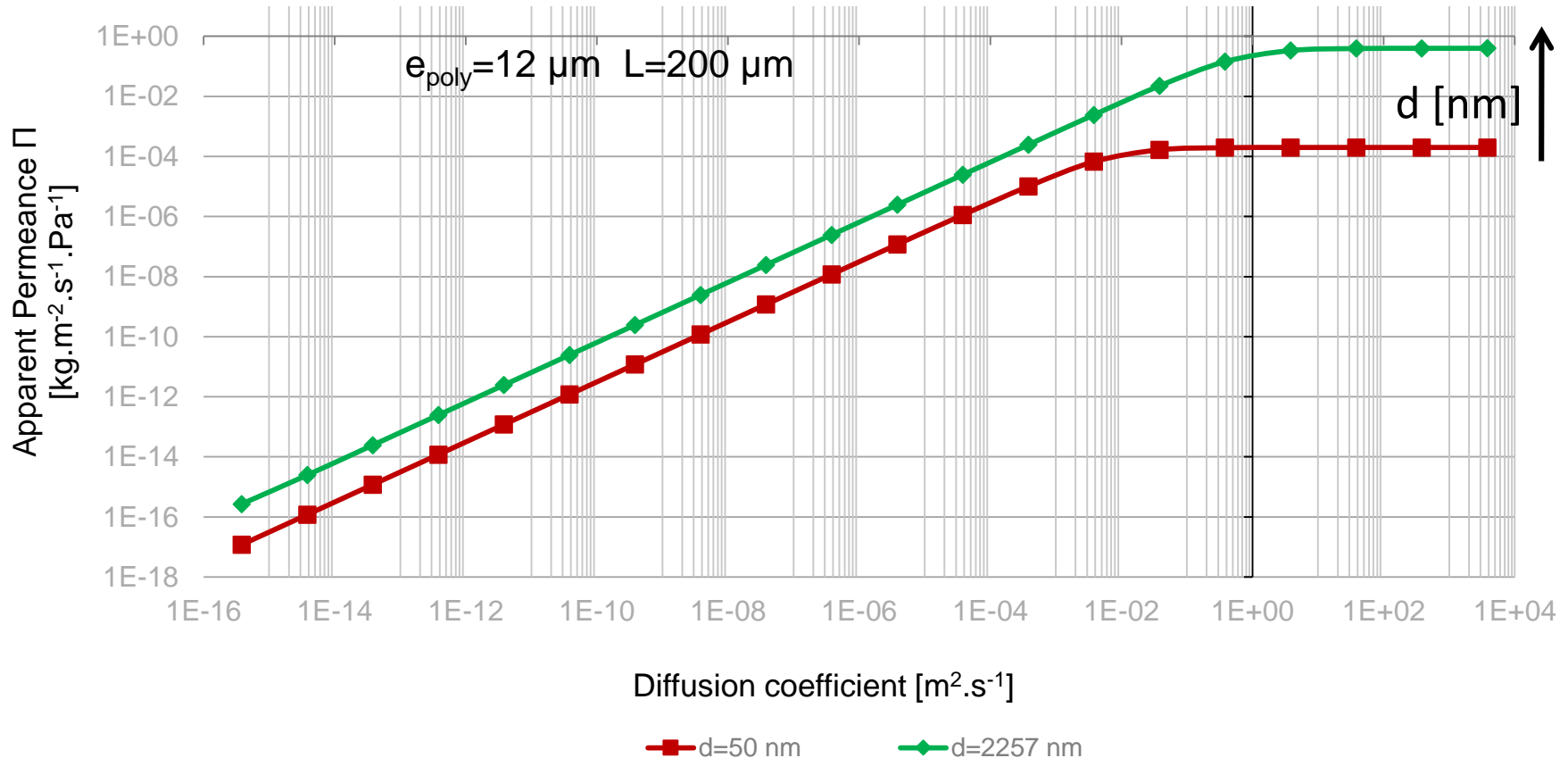
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# SOME RESULTS WITH PET M1F

- Diffusivity and solubility coefficients ( $D$ ,  $S$ )

Apparent permeance as function of the diffusion coefficient  
for different sizes of defects

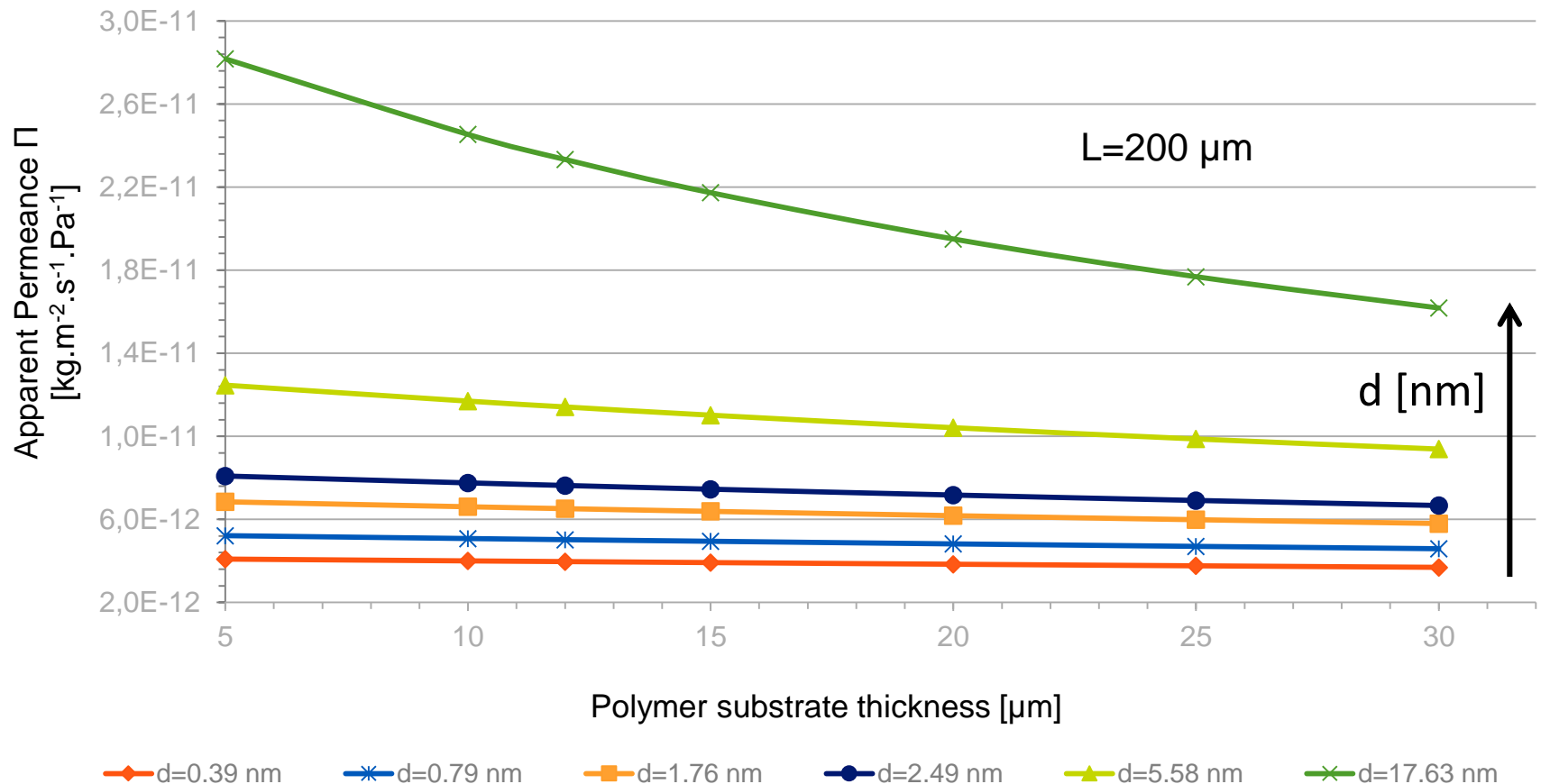




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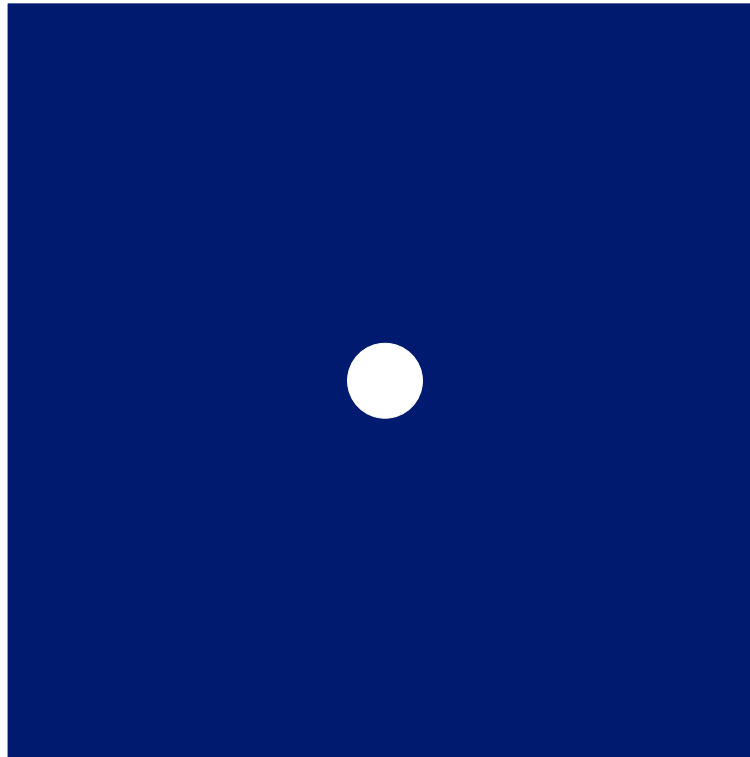
## ■ Polymer substrate thickness ( $e_{\text{poly}}$ )

Apparent permeance as function of the polymer substrate thickness  
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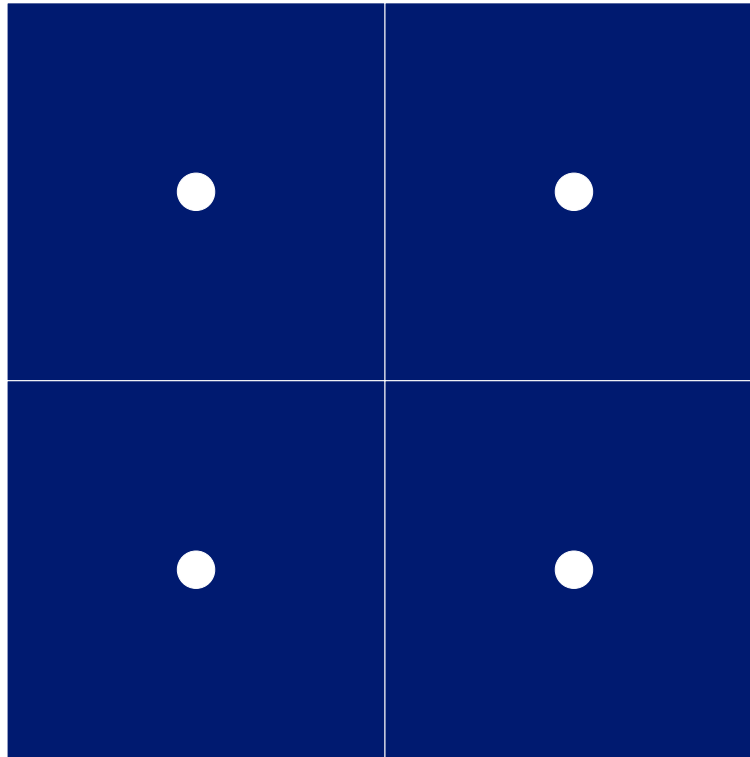
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- Defects size (diameter  $d$ ) and distance between the defects ( $L$ )



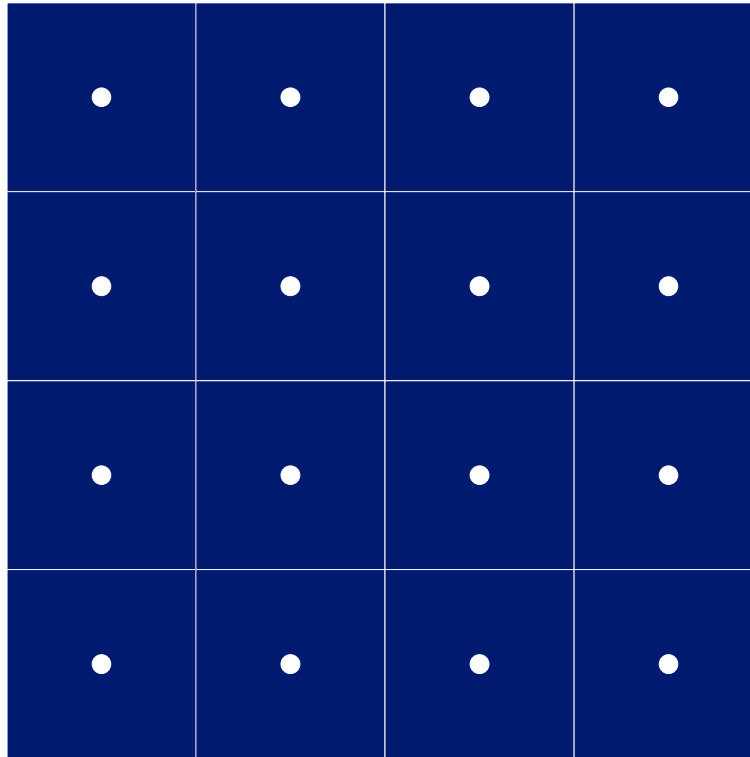
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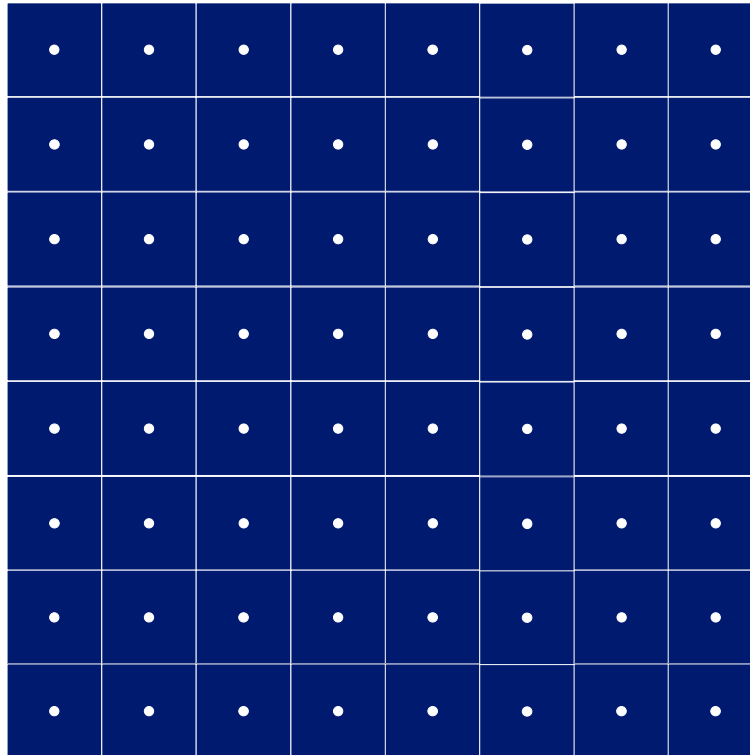
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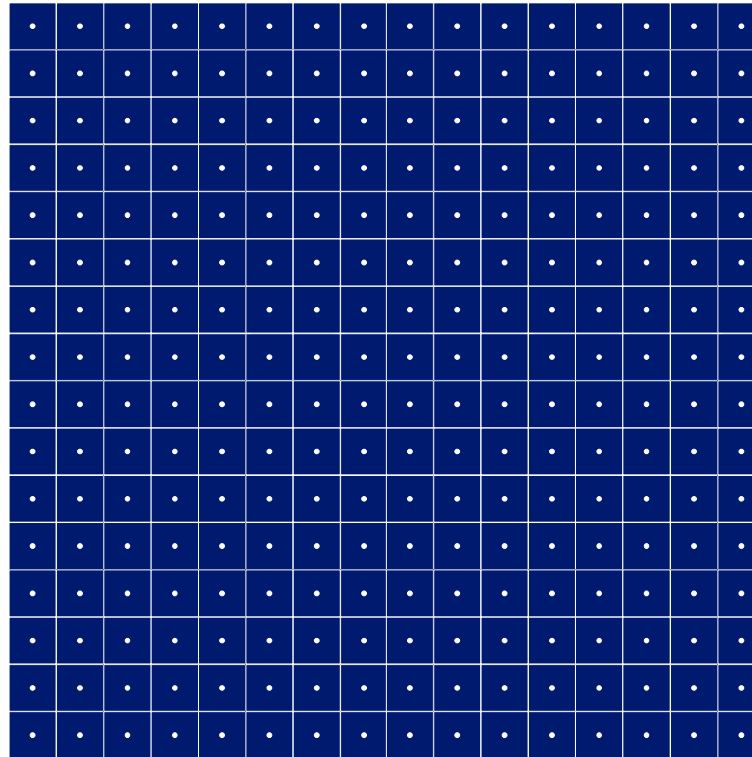
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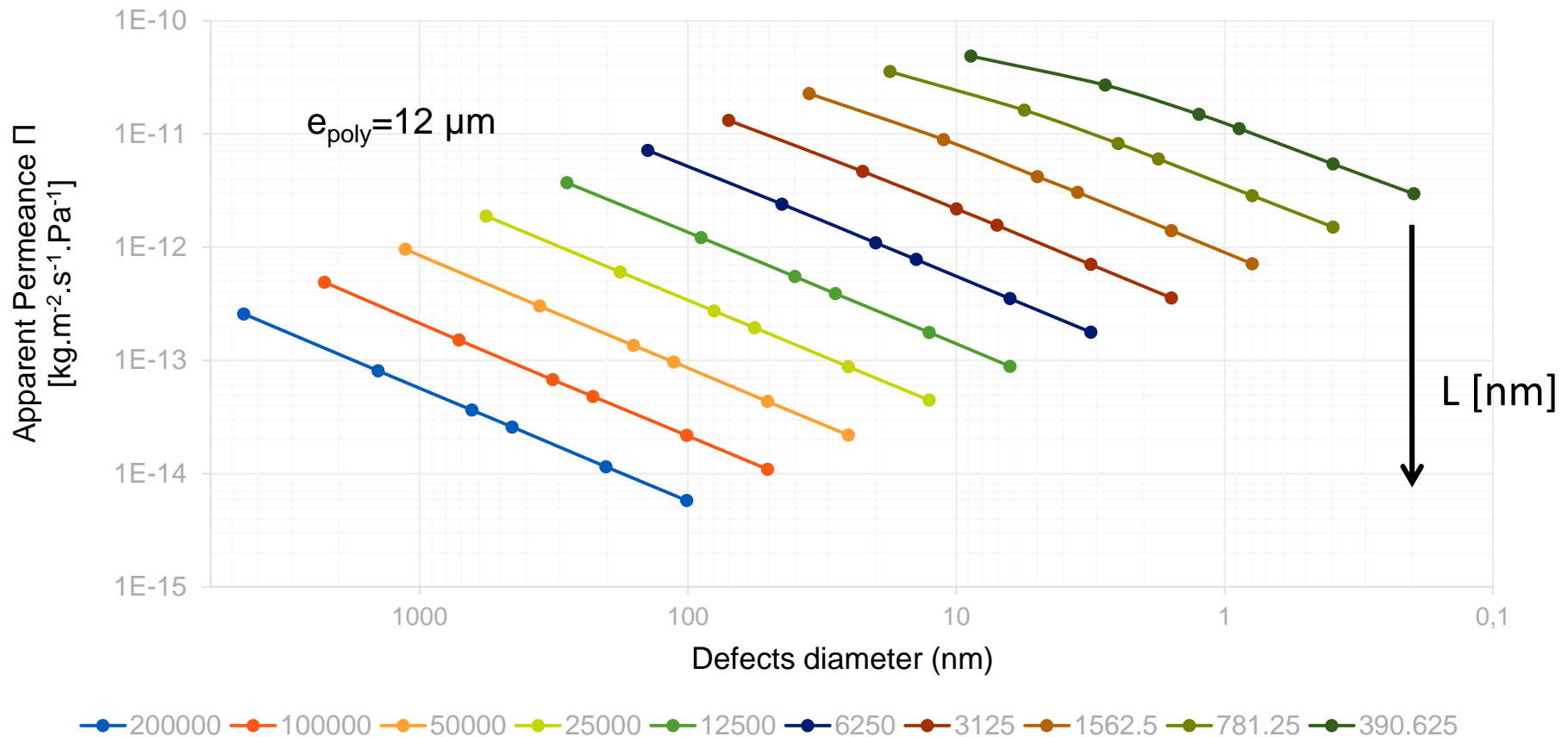
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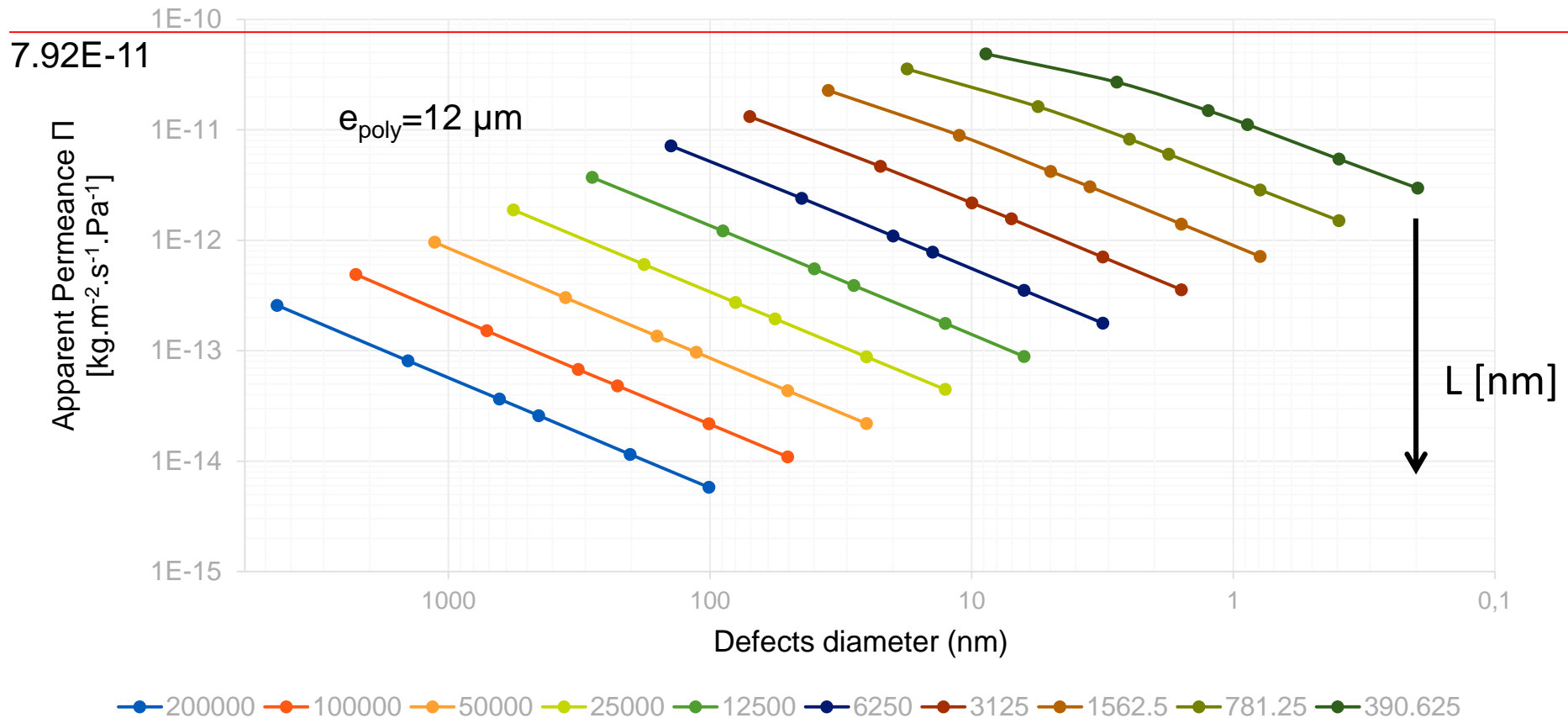
Apparent permeance as function of the defects size  
for different distances between defects (nm)



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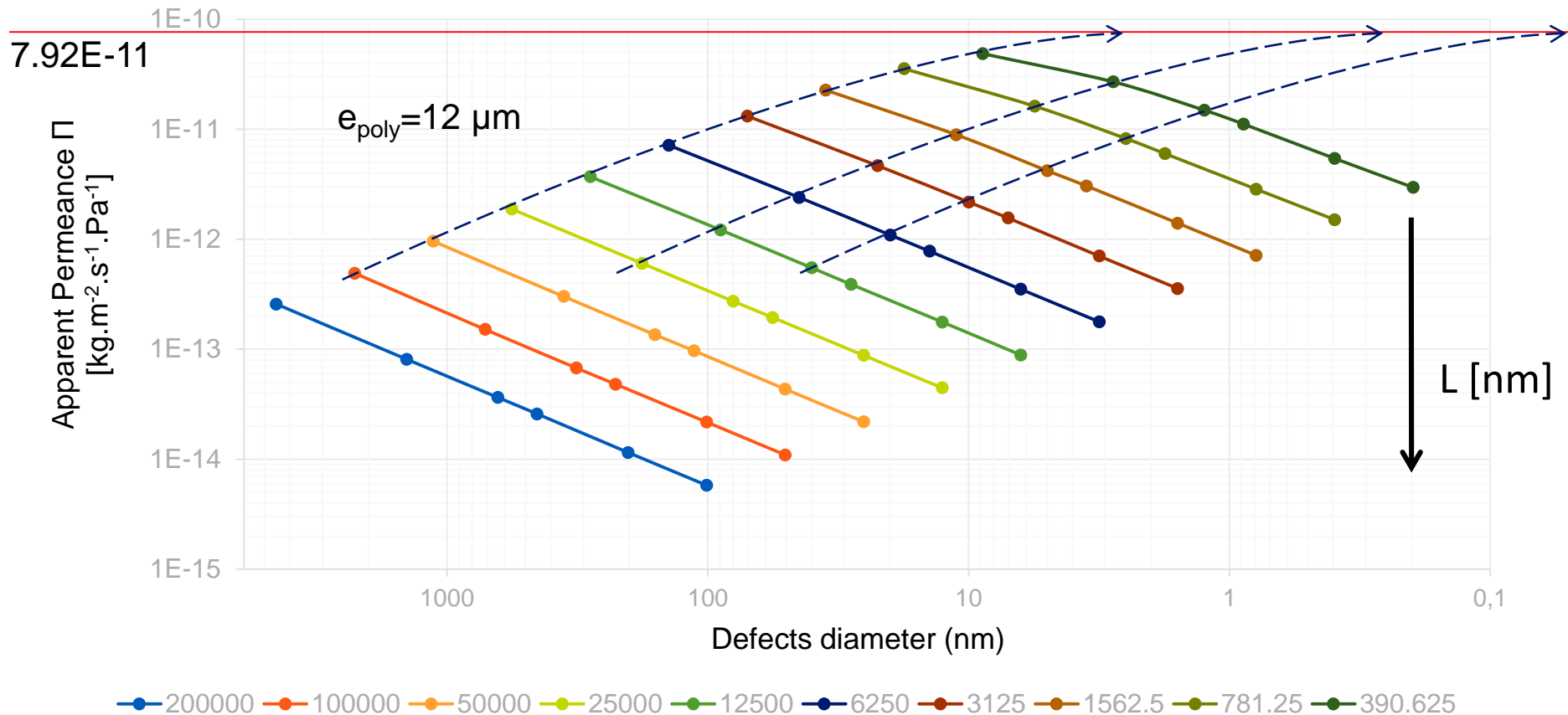




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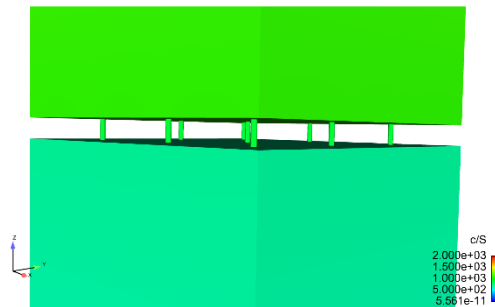
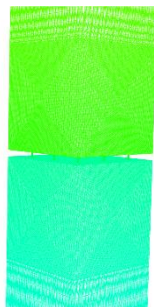




# CONCLUSION AND OUTLOOK

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- The gas transfer is not only regulated by the defects of the coating layer.
- The permeance gradually decreases with the thickness of the polymer substrate... and this effect tends to diminish when there are less defects.
- The permeance is more influenced by the distance between defects than the defects size.
- When the mean distance between defects is very small, the permeance of the PET M1F approaches that of the uncoated.



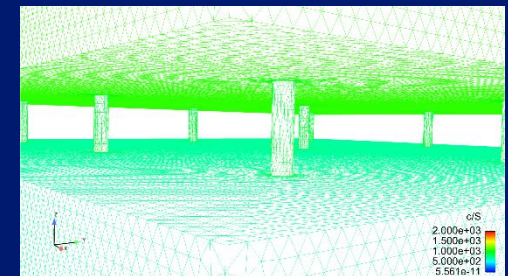
\* OUTLOOK \*

MULTILAYERS MODELLING

EXPERIMENTAL OBSERVATIONS

COMPARISON BETWEEN  
EXPERIMENTAL PERMEANCES  
AND CALCULATED PERMEANCES

RESEARCH ANALYTICAL MODEL



**Thanks for  
listening**